



US007923194B2

(12) **United States Patent**  
**Kohl et al.**

(10) **Patent No.:** **US 7,923,194 B2**  
(45) **Date of Patent:** **\*Apr. 12, 2011**

(54) **SACRIFICIAL COMPOSITIONS AND METHODS OF FABRICATING A STRUCTURE USING SACRIFICIAL COMPOSITIONS**

**G03F 7/20** (2006.01)

**G03F 7/26** (2006.01)

(52) **U.S. Cl.** ..... **430/270.1; 430/311; 430/330; 430/331; 430/913**

(75) Inventors: **Paul A. Kohl**, Atlanta, GA (US); **Paul J. Joseph**, Lawrenceville, GA (US); **Hollie K. Reed**, Pearland, TX (US); **Sue Ann Bidstrup-Allen**, Atlanta, GA (US); **Celesta E. White**, Katy, TX (US); **Clifford Henderson**, Douglasville, GA (US)

(58) **Field of Classification Search** ..... **430/270.1, 430/330, 311, 331, 913**  
See application file for complete search history.

(56) **References Cited**

**U.S. PATENT DOCUMENTS**

4,622,114	A	11/1986	Glass et al.
5,525,190	A	6/1996	Wojnarowski et al.
5,587,224	A	12/1996	Hsieh et al.
5,731,123	A	3/1998	Kawamura et al.
6,319,655	B1	11/2001	Wong et al.
6,610,593	B2	8/2003	Kohl et al.
6,743,570	B2	6/2004	Harnett et al.
6,788,867	B2	9/2004	Mule et al.
7,459,267	B2 *	12/2008	Kohl et al. .... 430/330
2002/0127768	A1	9/2002	Badir et al.
2002/0136481	A1	9/2002	Mule et al.
2004/0132855	A1	7/2004	Kohl et al.
2004/0137728	A1	7/2004	Gallagher et al.
2004/0248034	A1	12/2004	Henderson et al.

(73) Assignee: **Georgia Tech Research Corporation**, Atlanta, GA (US)

(\*) Notice: Subject to any disclaimer, the term of this patent is extended or adjusted under 35 U.S.C. 154(b) by 98 days.

This patent is subject to a terminal disclaimer.

(21) Appl. No.: **12/273,114**

(22) Filed: **Nov. 18, 2008**

(65) **Prior Publication Data**

US 2009/0107952 A1 Apr. 30, 2009

**Related U.S. Application Data**

(62) Division of application No. 11/398,183, filed on Apr. 5, 2006, now Pat. No. 7,459,267, which is a division of application No. 10/699,330, filed on Oct. 31, 2003, now Pat. No. 7,052,821.

(60) Provisional application No. 60/423,013, filed on Nov. 1, 2002.

(51) **Int. Cl.**

**G03F 7/00** (2006.01)

**G03F 7/004** (2006.01)

**OTHER PUBLICATIONS**

P. A. Kohl, Q. Zhao, K. S. Patel, D. Schmidt, S. A. Bidstrup-Allen, R. Shick, S. Jayaraman, Air-Gaps for Electrical Interconnections, Electrochemical and Solid State Lett, vol. 1, p. 49, 1998. cited by other.

(Continued)

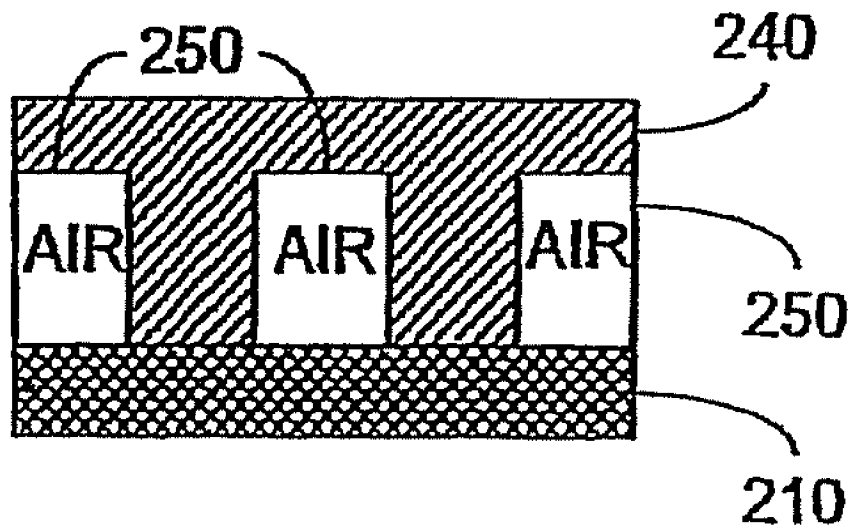
*Primary Examiner* — Amanda C. Walke

(74) *Attorney, Agent, or Firm* — Bernard Berman

(57) **ABSTRACT**

Compositions, methods of use thereof, and methods of decomposition thereof, are provided. One exemplary composition, among others, includes a polymer and a catalytic amount of a negative tone photoinitiator.

**7 Claims, 21 Drawing Sheets**

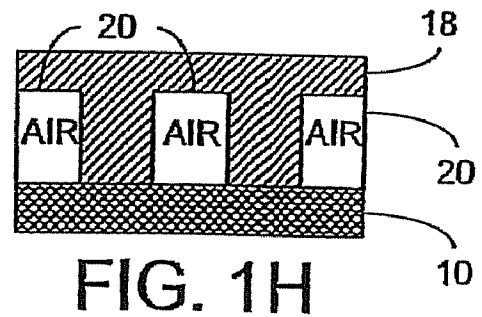
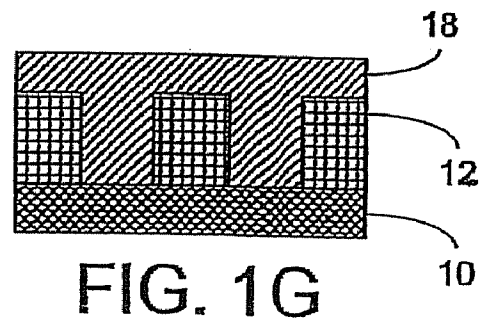
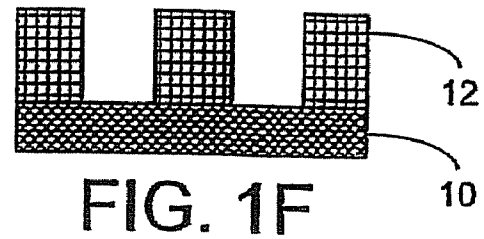
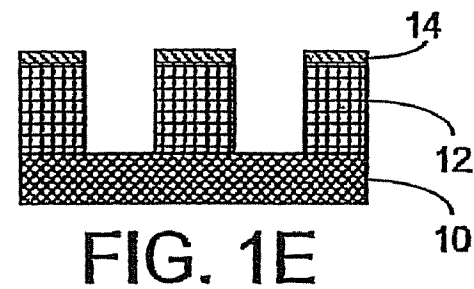
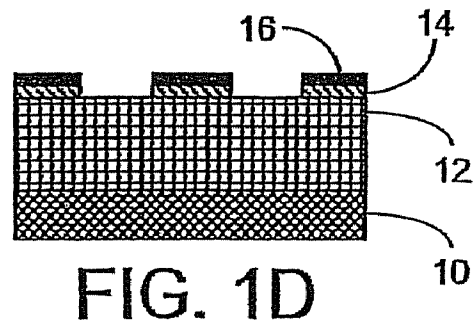
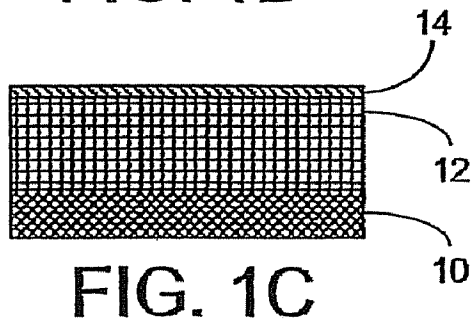
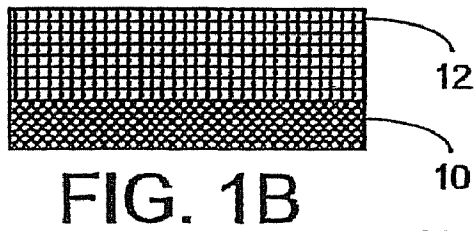
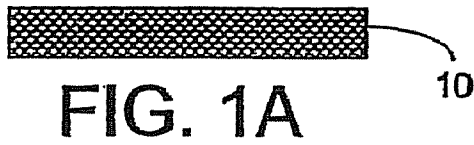


## OTHER PUBLICATIONS

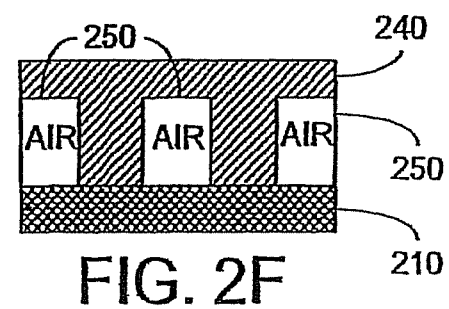
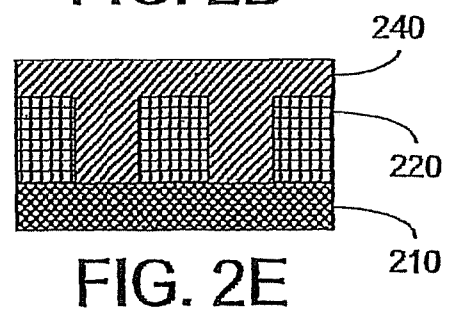
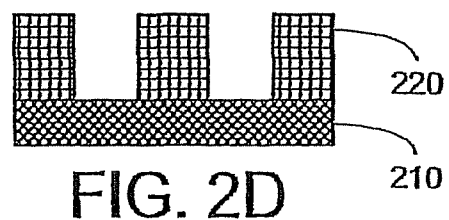
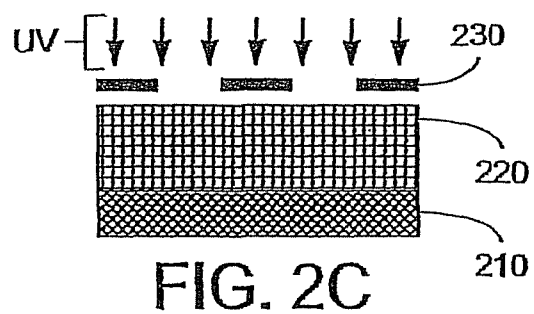
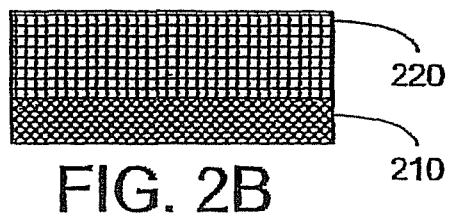
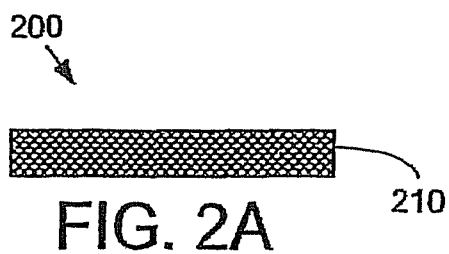
- P. A. Kohl, D. M. Bhusari, M. Wedlake, C. Case, B. C. Lee, R. J. Gutmann, R. Shick, Air-Gaps in 0.3  $\mu\text{m}$  Electrical Interconnections, *IEEE Electron Device Lett.*, vol. 21, p. 557, 2000. cited by other.
- H. A. Reed, C. E. White, V. Rao, S. A. Bidstrup-Allen, C. L. Henderson, P. A. Kohl, Fabrication of microchannels using polycarbonates as sacrificial materials, *J Micromech Microeng.*, vol. 11(6), p. 733, 2001. cited by other.
- D. Bhusari, H. A. Reed, M. Wedlake, A. Padovani, S. A. Bidstrup-Allen, P. A. Kohl, Fabrication of Air-Channel Structures for Microfluidic, Microelectromechanical, and Microelectronic Applications, *J Micromech. Microeng.*, vol. 10 (3), p. 400, 2001. cited by other.
- M. B. Anand, M. Yamada, H. Shibata, Use of Gas as Low-k Interlayer Dielectric in LSI's: Demonstration of Feasibility, *IEEE Transactions on Electron Devices*, vol. 44, p. 1965, 1997. cited by other.
- C. K. Harriott, G. W. Coates, H. G. Craighead, Heat-depolymerizable polycarbonates as electron beam patternable sacrificial layers for nanofluidics, *J Vac. Sci. Technol B.*, vol. 19(6), p. 2842, 2001. cited by other.
- L. S. Lee, K. K. Gleason, Hot Filament Chemical Vapor Deposition of Polyoxymethylene as a Sacrificial Layer for Fabricating Air Gaps, *Electrochemical and Solid State Lett.*, vol. 4, p. G81, 2001. cited by other.
- H.-J. Suh, P. Bharathi, D. J. Beebe, J. S. Moore, Dendritic Material as a Dry-Release Sacrificial Layer, *J. Microelectromech. Syst.*, vol. 9(2), pp. 198-205, 2000. cited by other.
- J. P. Gravesen, J. Bianejer, O. S. Jensen, Microfluidics—A Review, *J Micromech. Microeng.*, vol. 3, p. 168, 1993. cited by other.
- R. F. Service, Microchip Arrays Put DNA on the Spot, *Science*, vol. 282, p. 396, 1998. cited by other.
- J. V. Crivello, J. H. W. Lam, Diaryliodonium Salts. A New Class of Photoinitiators for Cationic Polymerization, *Macromolecules*, vol. 10(6), p. 1307, 1977. cited by other.
- R. Taylor, The Nature of the Transition State in Ester Pyrolysis. Part II. The Relative Rates of Pyrolysis of Ethyl, Isopropyl, and t-Butyl Acetates, Phenylacetates, Benzoates, Phenyl Carbonates, and N-Phenylcarbamates, *J Chem. Soc. Chem. Commun.*, p. 1025, 1975. cited by other.
- S. Inoue, T. Tsuruta, T. Takada, N. Miyazaki, M. Kambe, T. Takaoka, Synthesis and Thermal Degradation of Carbon Dioxide-Epoxy Copolymer, *Appl. Polym. Symp.*, vol. 26, p. 257, 1975. cited by other.
- J. M. J. Frechet, F. Bouchard, F. M. Houlihan, E. Eichler, B. Kryczka, C. G. Wilson, Design and synthesis of novel allylic and benzylic copolycarbonates susceptible to acidolytic or thermolytic depolymerization, *Makromol. Chem. Rapid. Commun.*, vol. 7, p. 121, 1986. cited by other.
- S. C. Narang, S. T. Attarwala, Chemical Amplification in t-Diol Polycarbonate Resists, *Polym. Prepr.*, (Am. Chem. Soc. Div. Polym., Chem.), vol. 26(2), p. 323, 1985. cited by other.
- D. J. Drensbourg, J. R. Wildeson, J. C. Yarbrough, J. H. Reibenspies, Bis 2,6-difluorophenoxide Dimeric Complexes of Zinc and Cadmium and Their Phosphine Adducts: Lessons Learned Relative to Carbon Dioxide/Cyclohexene Oxide Alternating Copolymerization Processes Catalyzed by Zinc Phenoxides, *J Amer. Chem. Soc.*, vol. 122, p. 12487, 2000. cited by other.
- M. Murayama, F. Sanda, T. Endo, Anionic Ring-Opening Polymerization of a Cyclic Carbonate Having a Norbornene Structure with Amine Initiators, *Macromolecules*, vol. 31, p. 919, 1998. cited by other.
- Y. Toba, M. Saito, Y. Usui, Cationic Photopolymerization of Epoxides by Direct and Sensitized Photolysis of Onium Tetrakis(pentafluorophenyl)borate Initiators, *Macromolecules*, vol. 32(10), p. 3209, 1999. cited by other.
- J. V. Crivello, J. Lockhart, J. Lee, Diaryliodonium Salts as Thermal Initiators of Cationic Polymerization, *J. Polym. Sci Part A: Polym. Chem.*, vol. 21, p. 97, 1983. cited by other.

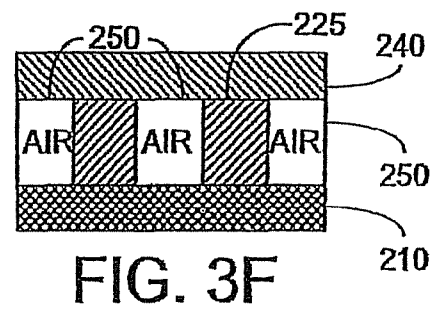
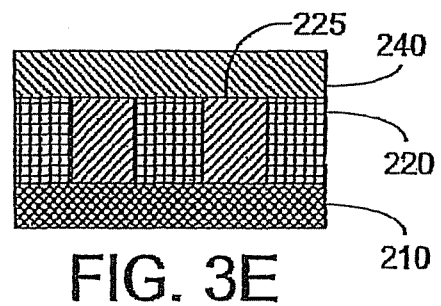
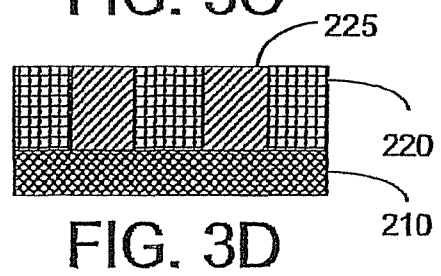
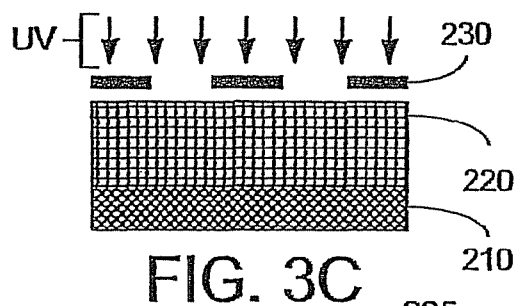
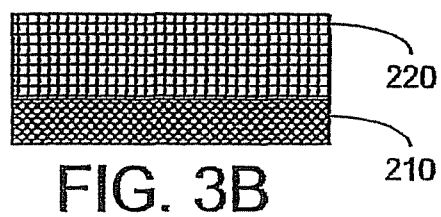
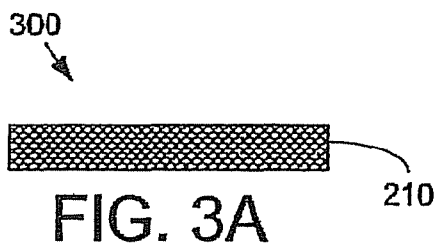
\* cited by examiner

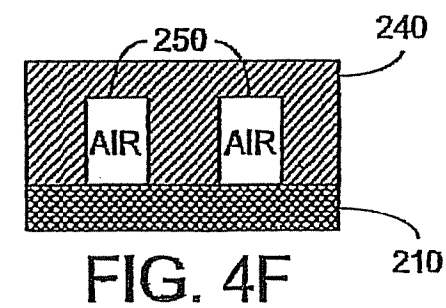
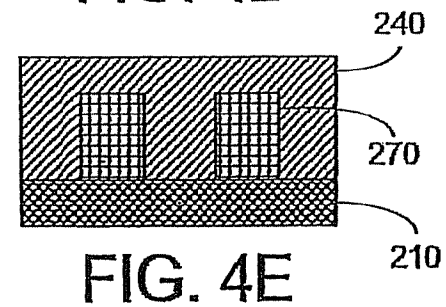
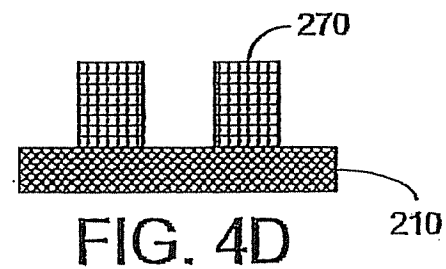
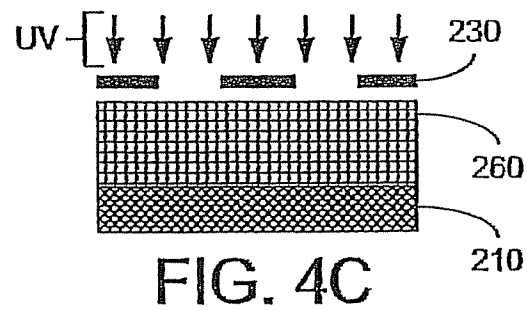
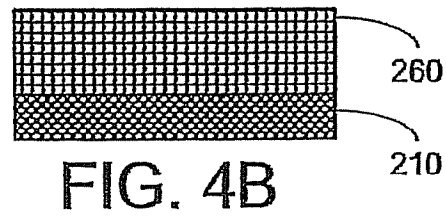
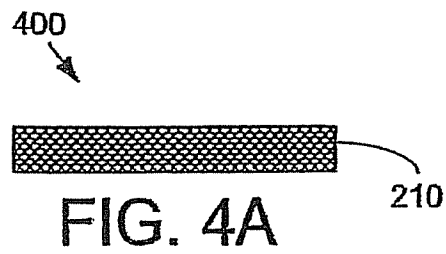
100  
↙



PRIOR ART







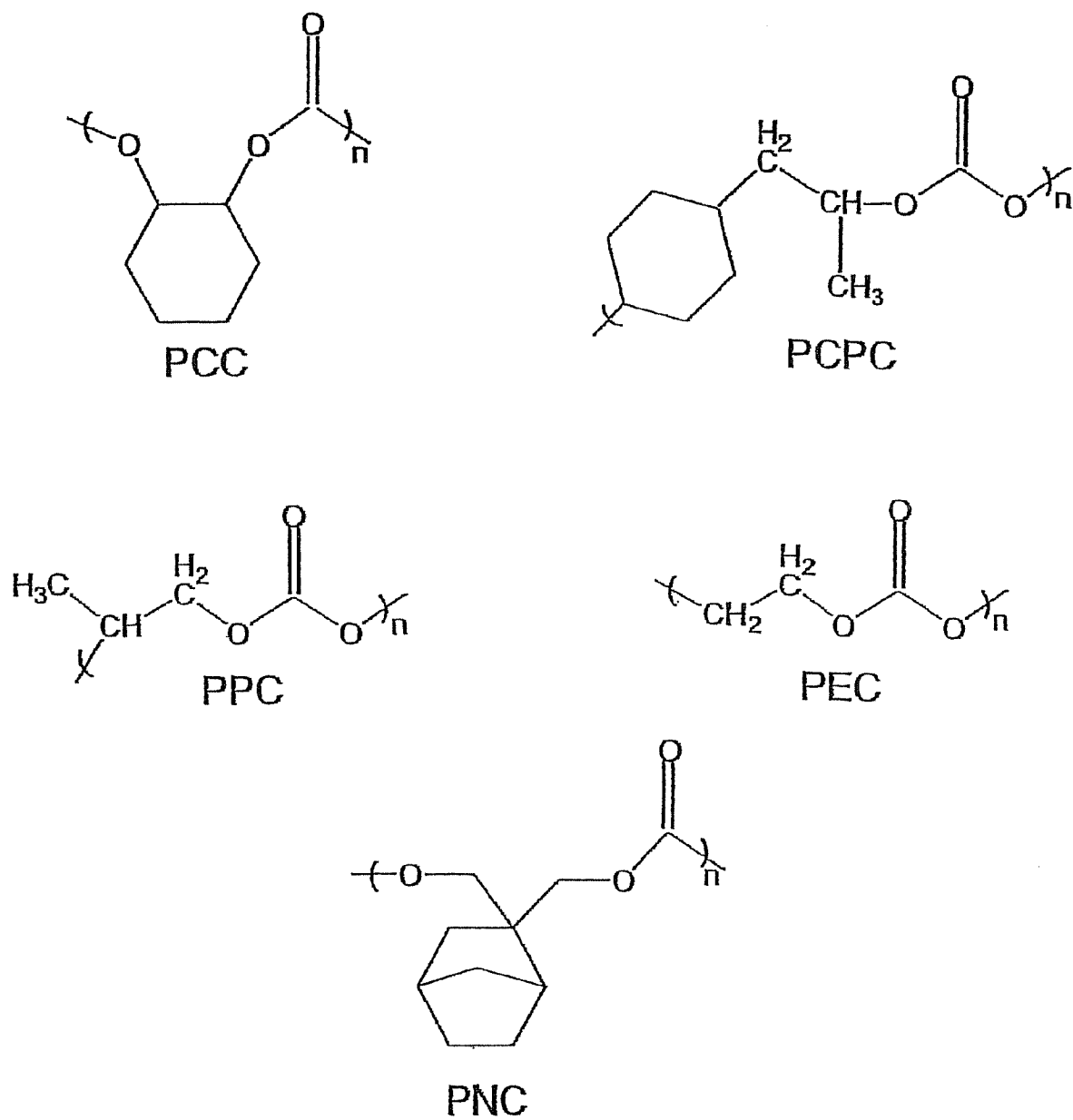
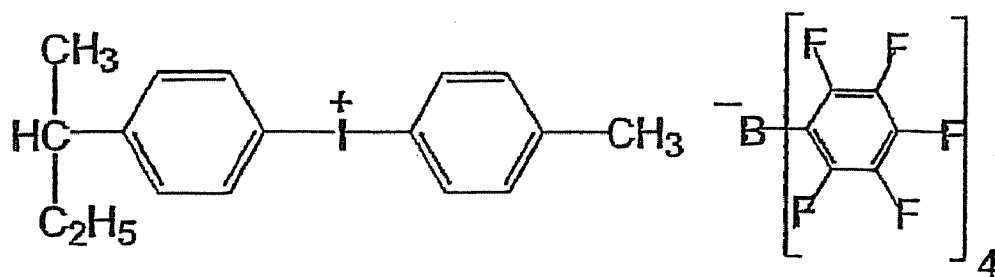
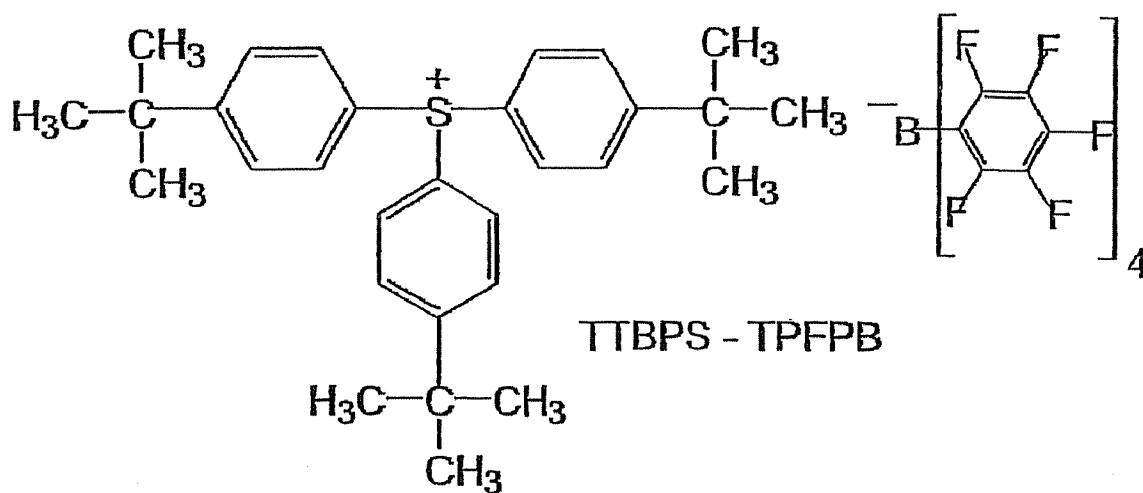


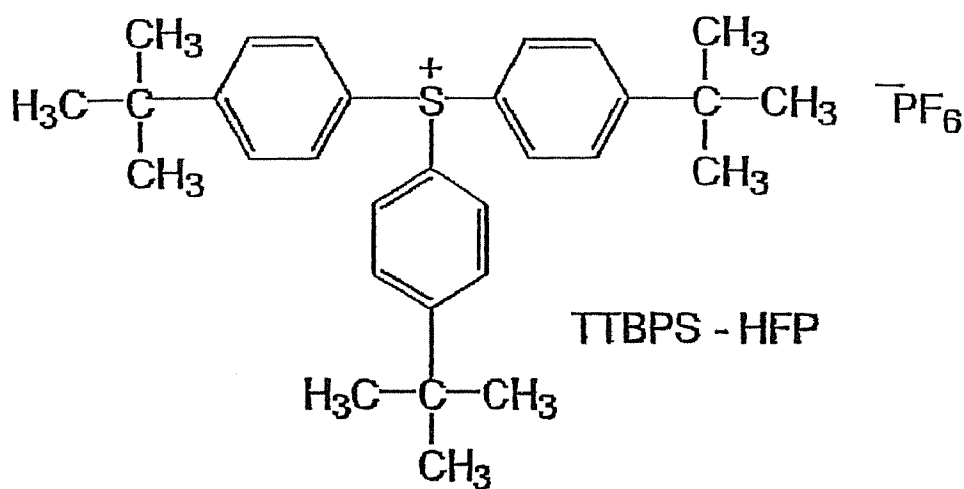
FIG. 5



DPI - TPFPB



TTBPS - TPFPB



TTBPS - HFP

FIG. 6



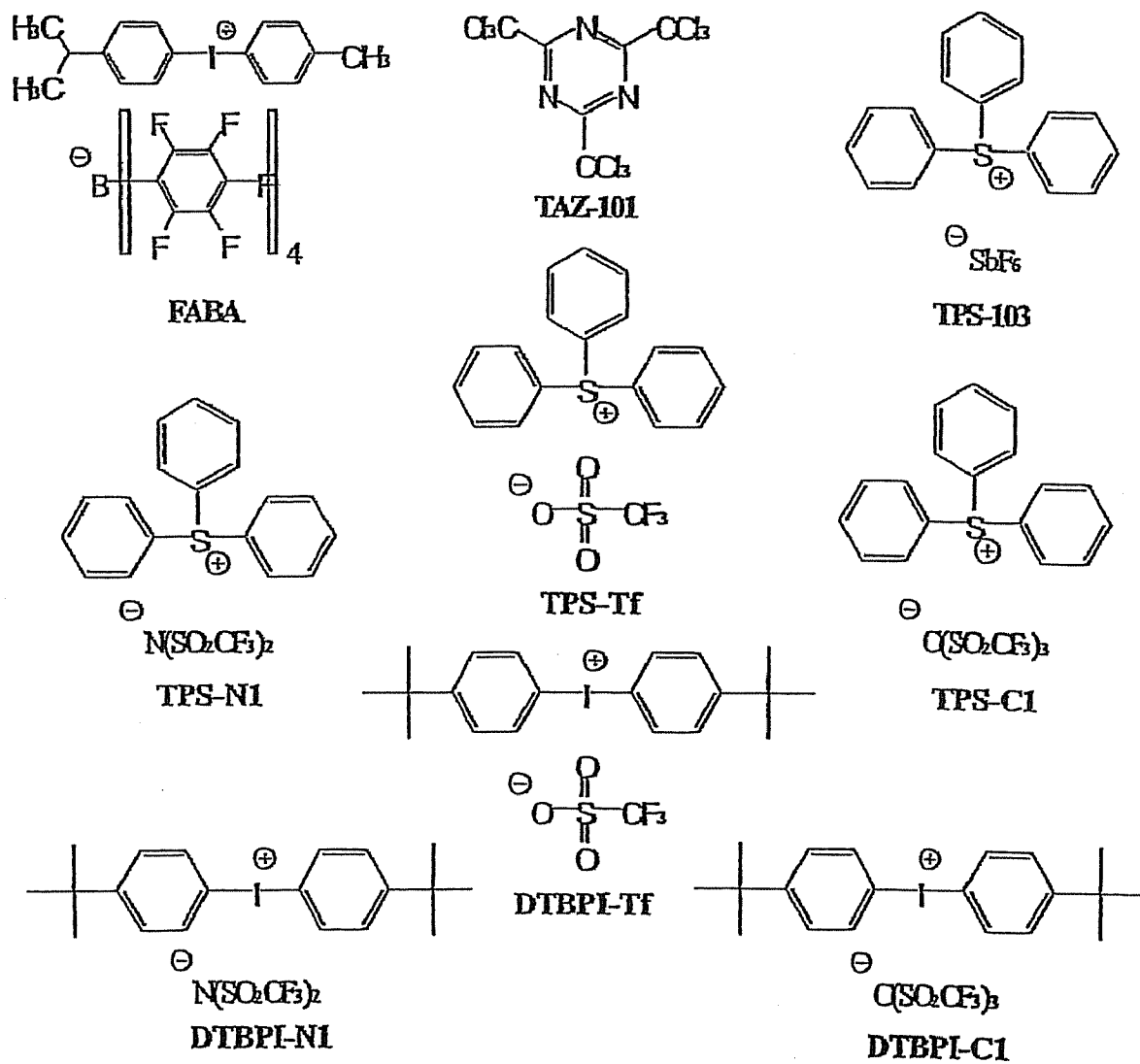


FIG. 7

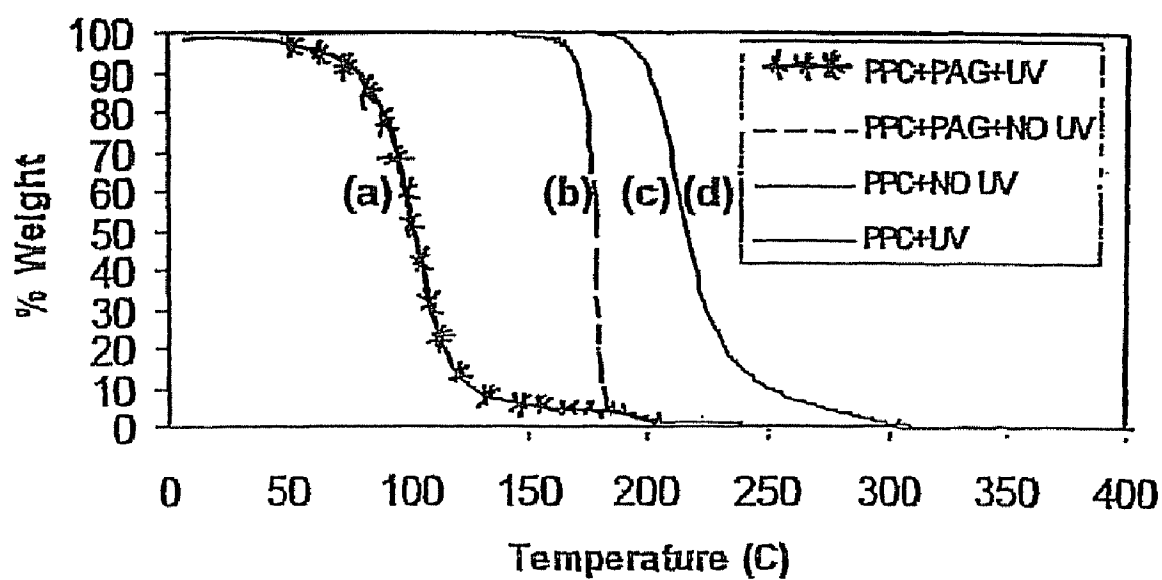


FIG. 8

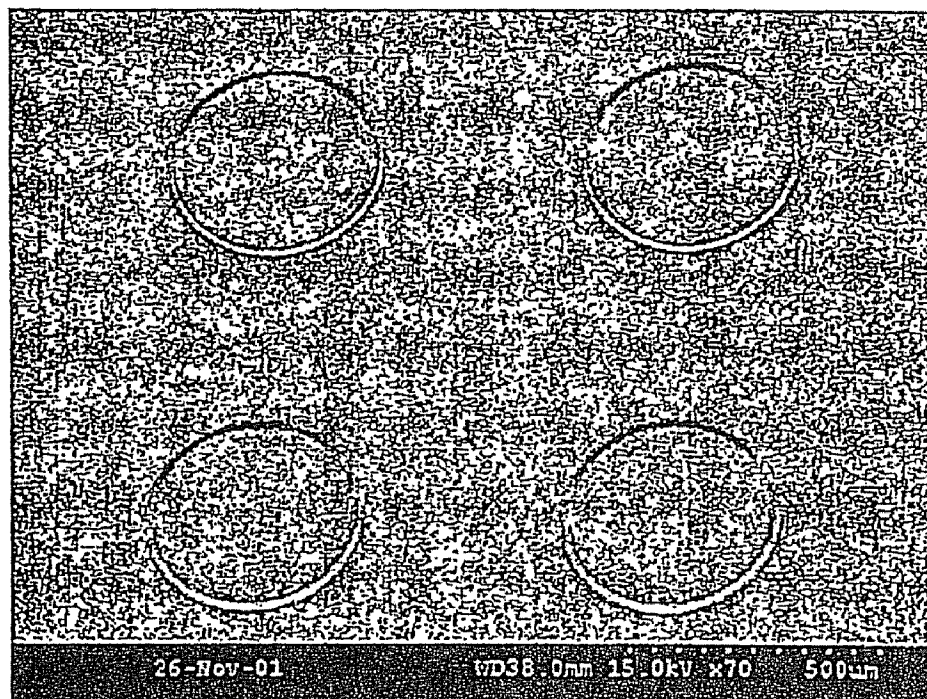
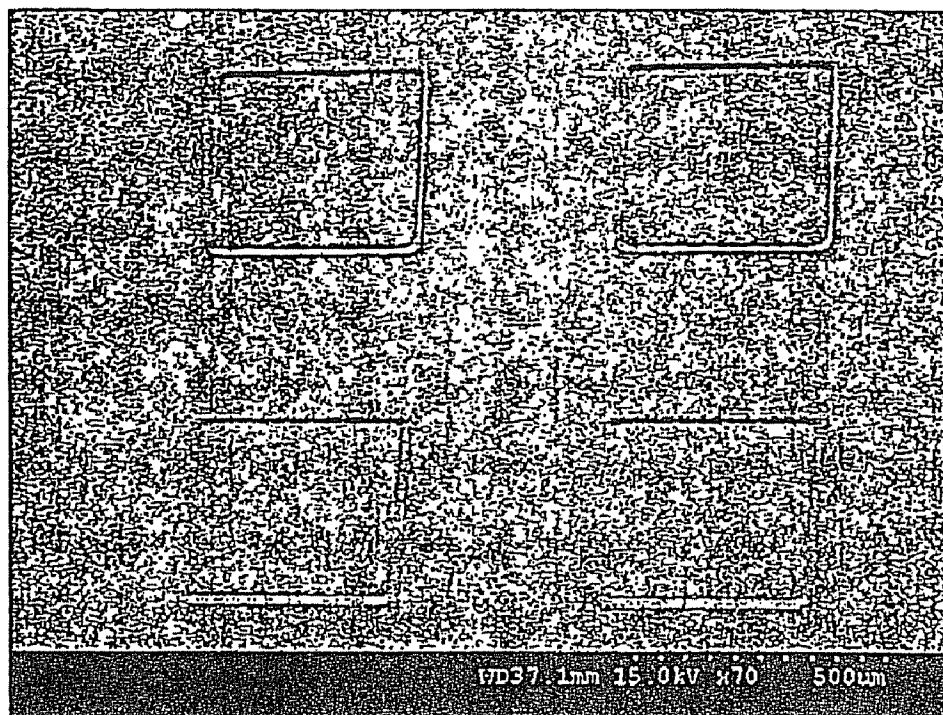


FIG. 9

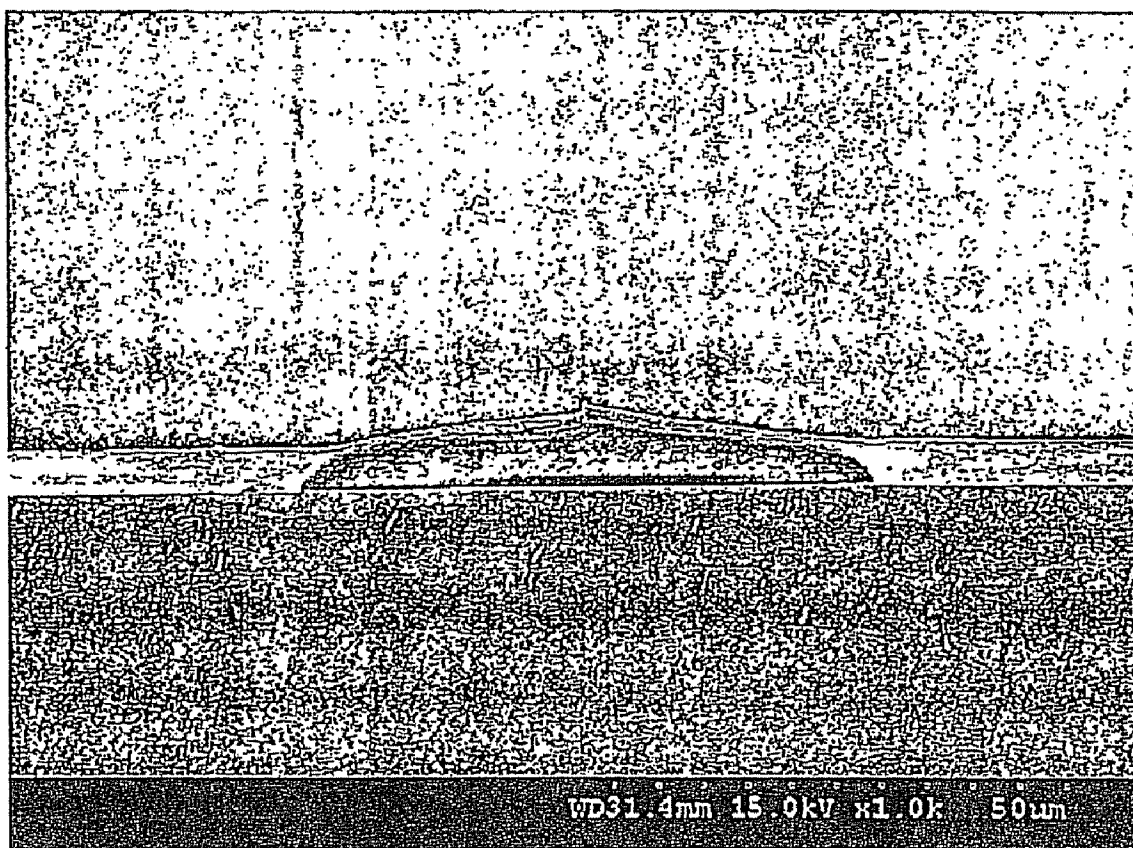


FIG. 10

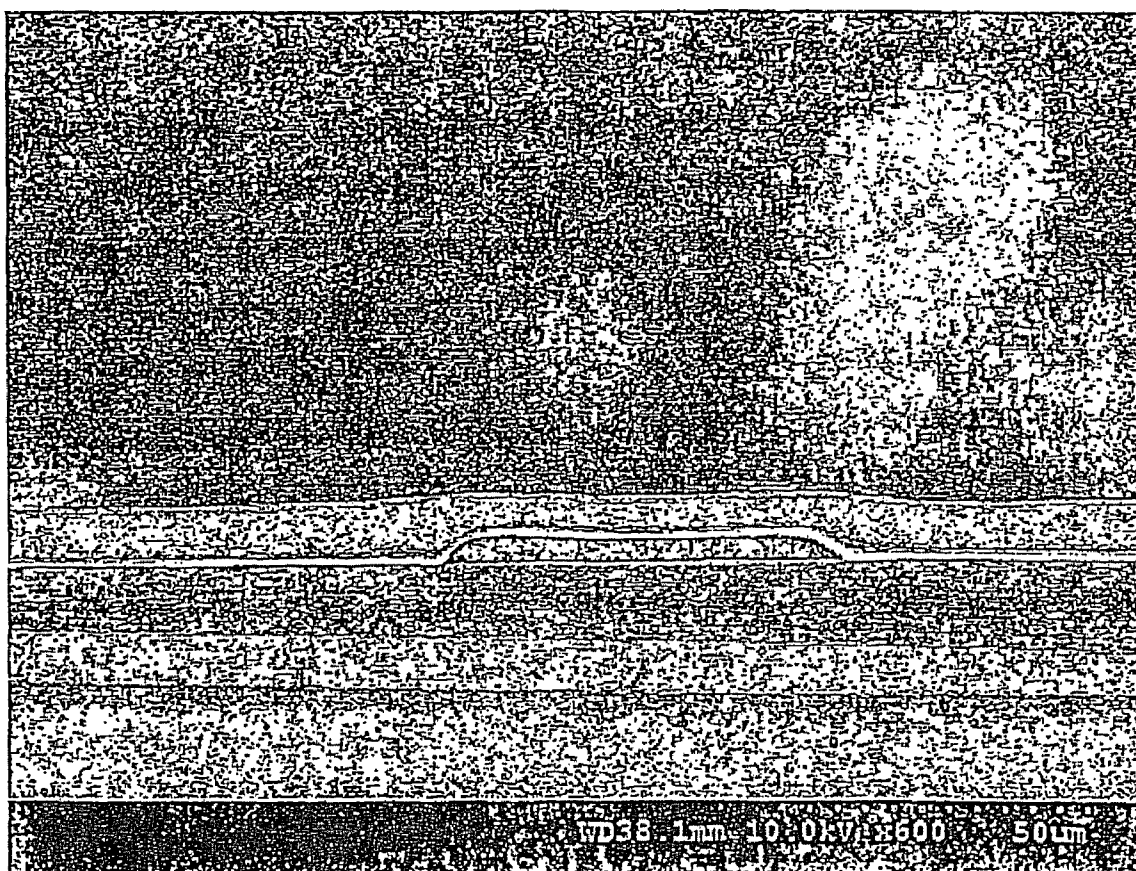


FIG. 11

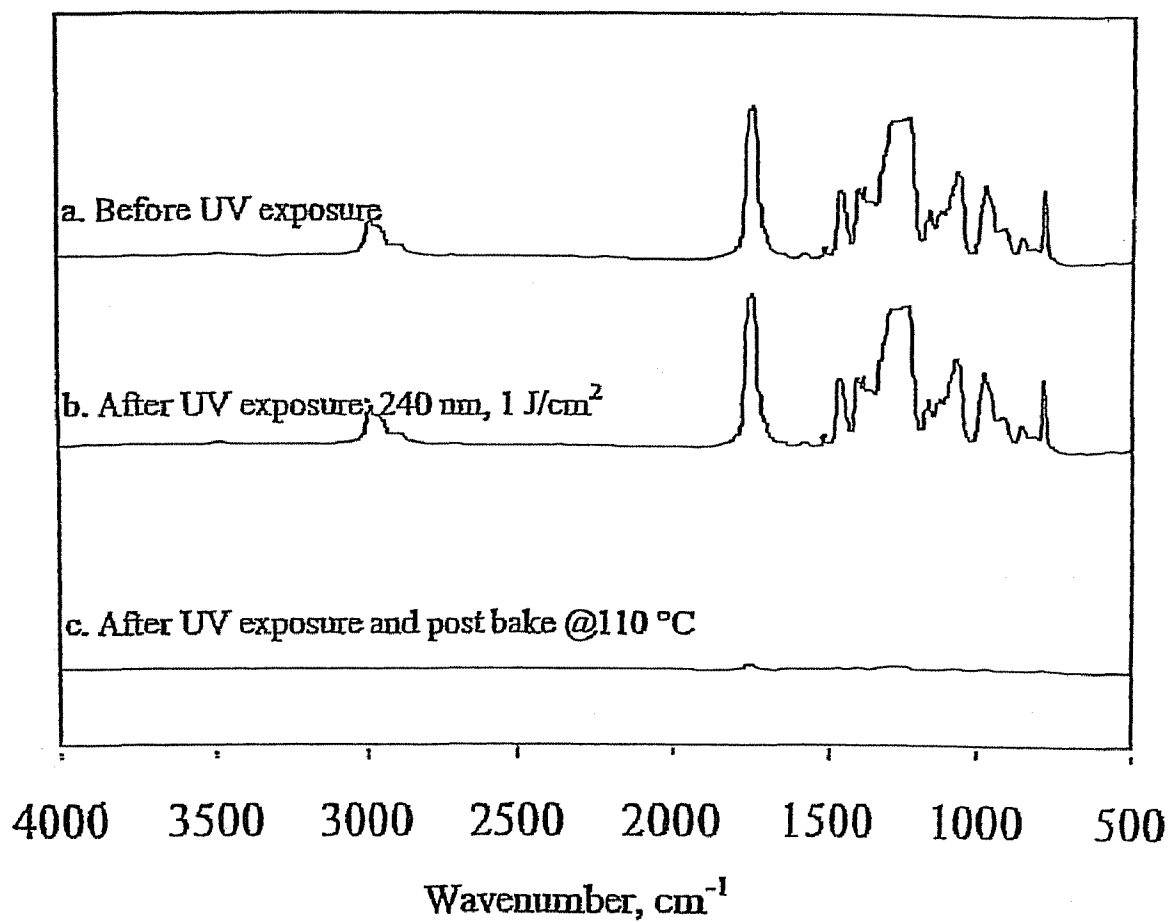


FIG. 12

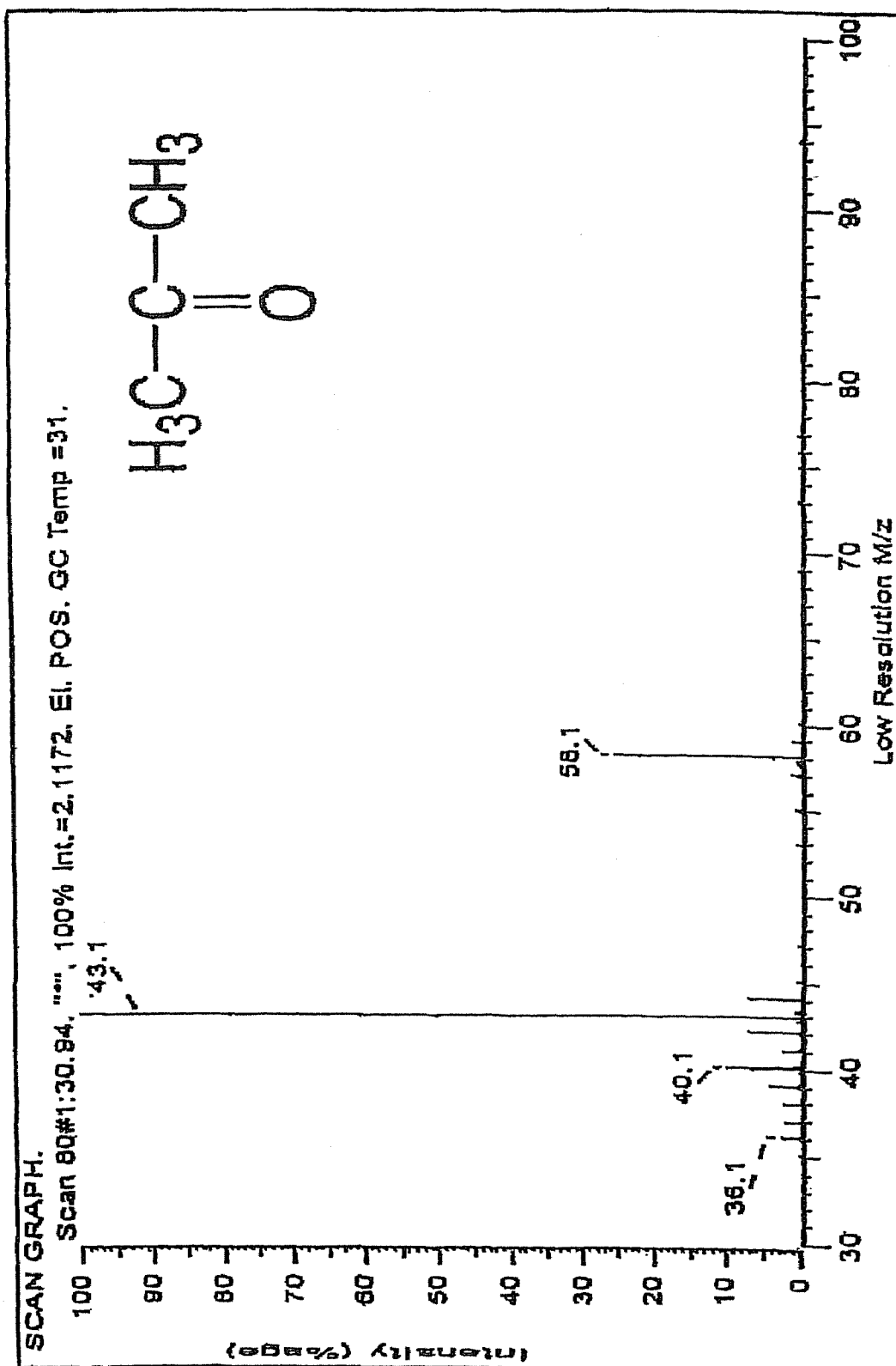


FIG. 13A

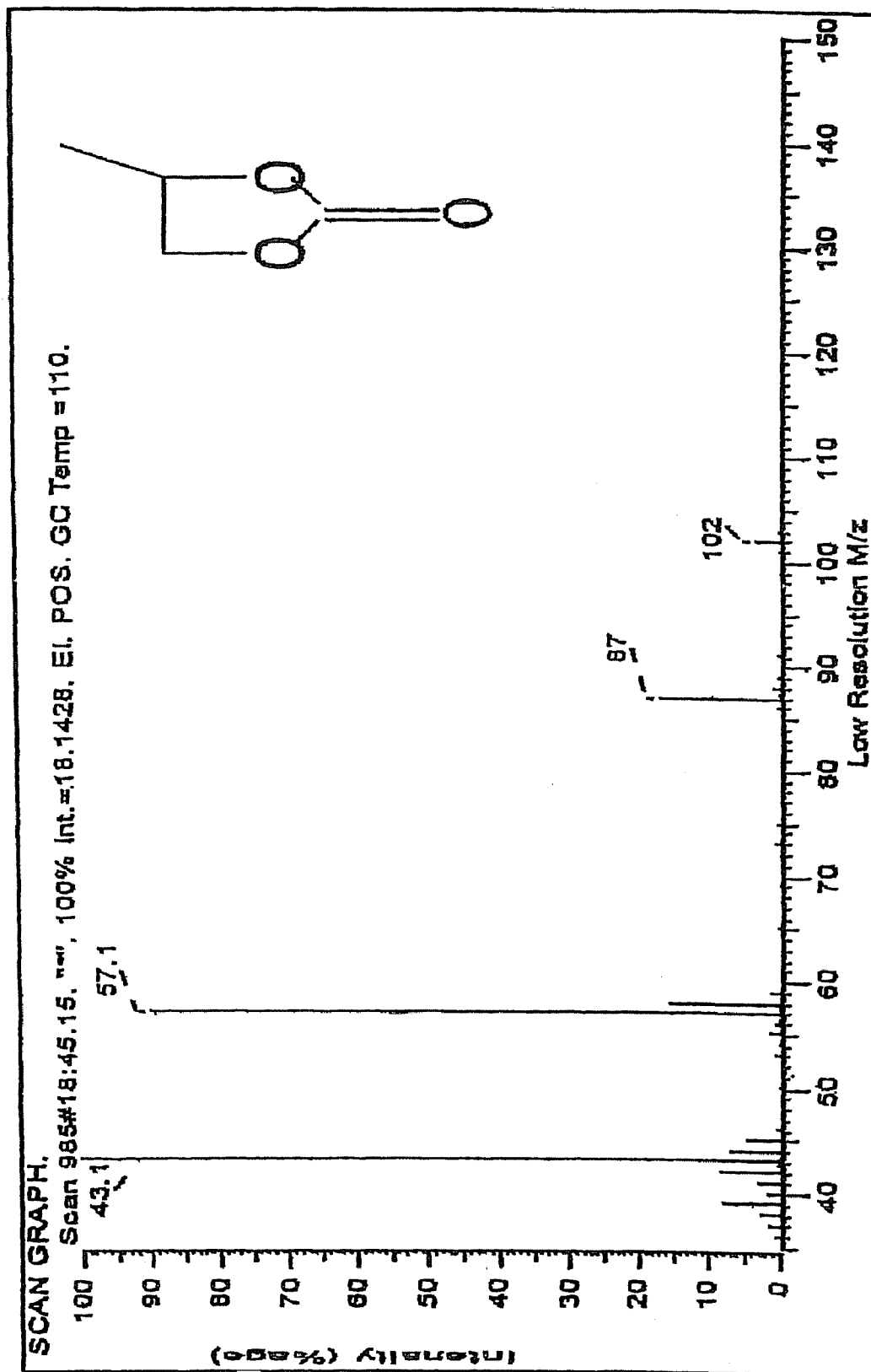


FIG. 13B



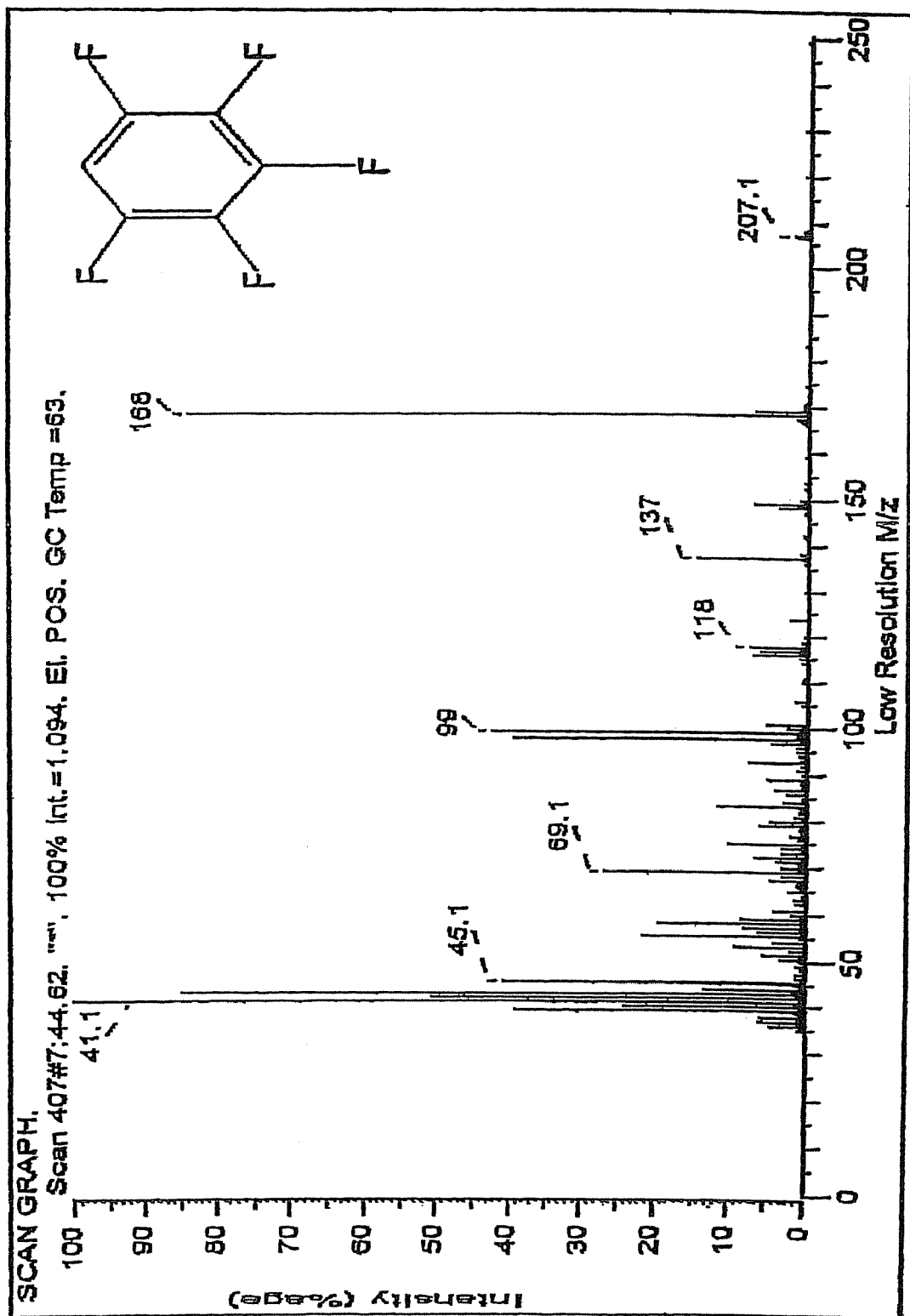


FIG. 13C

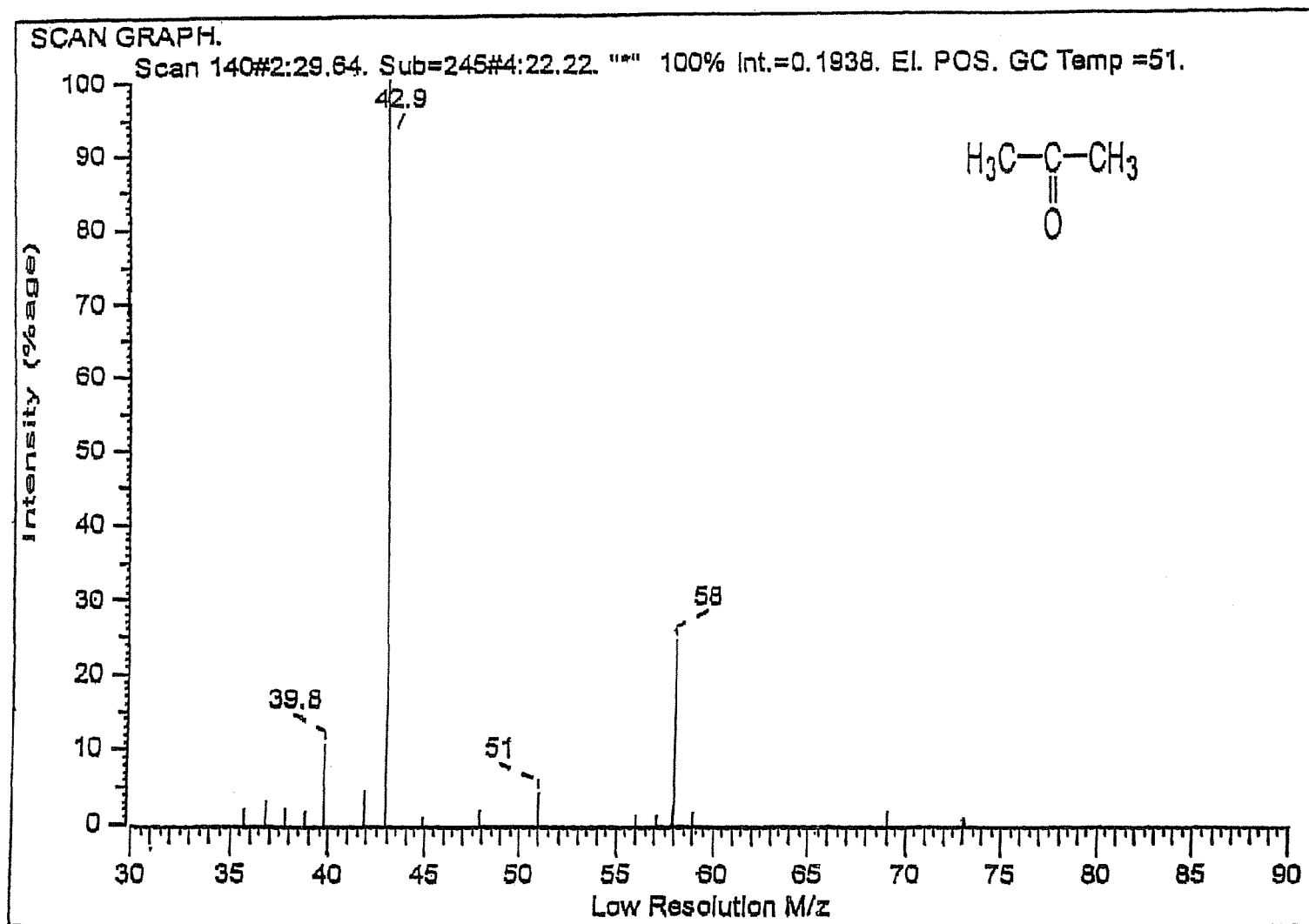


FIG. 14A

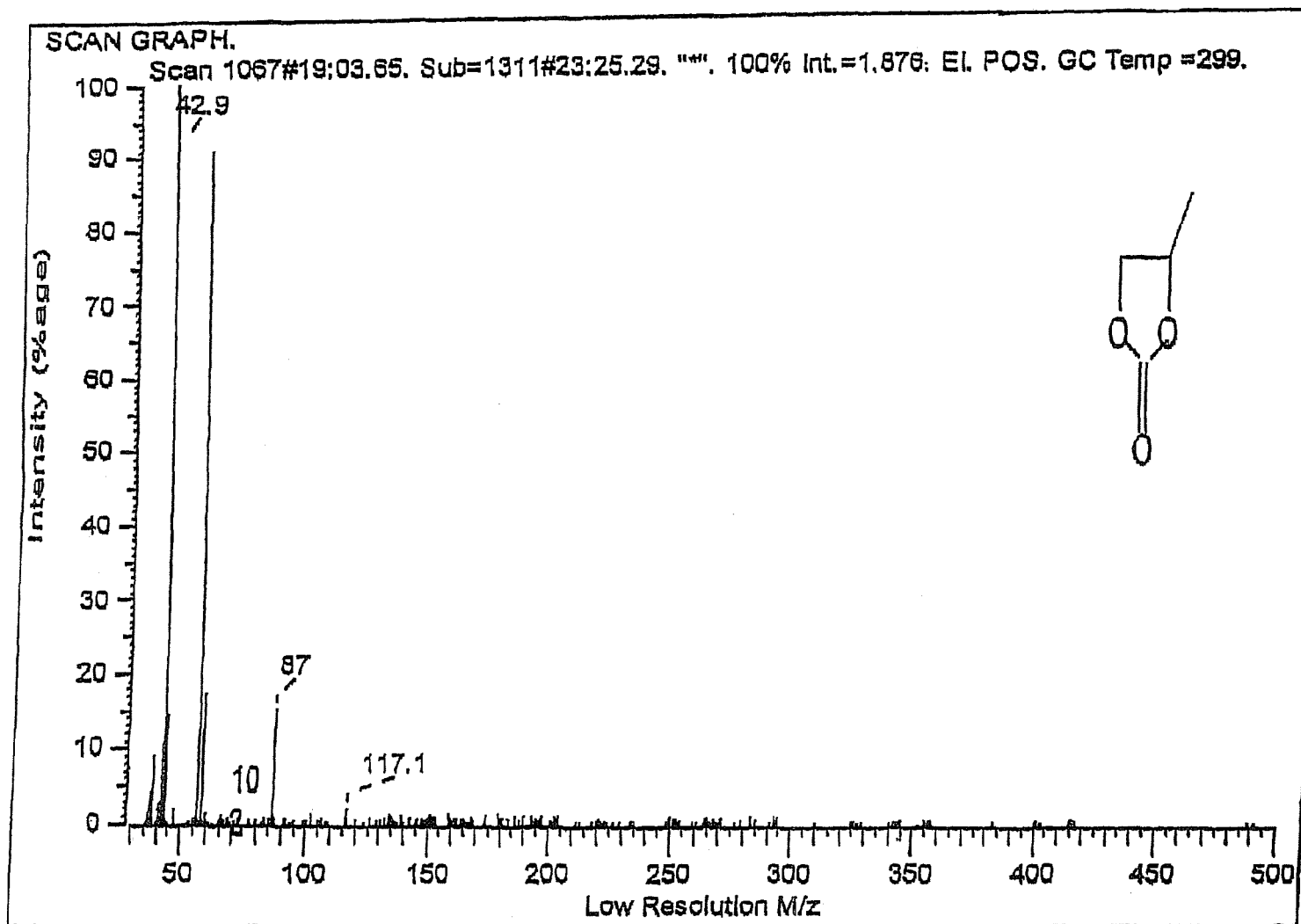


FIG. 14B

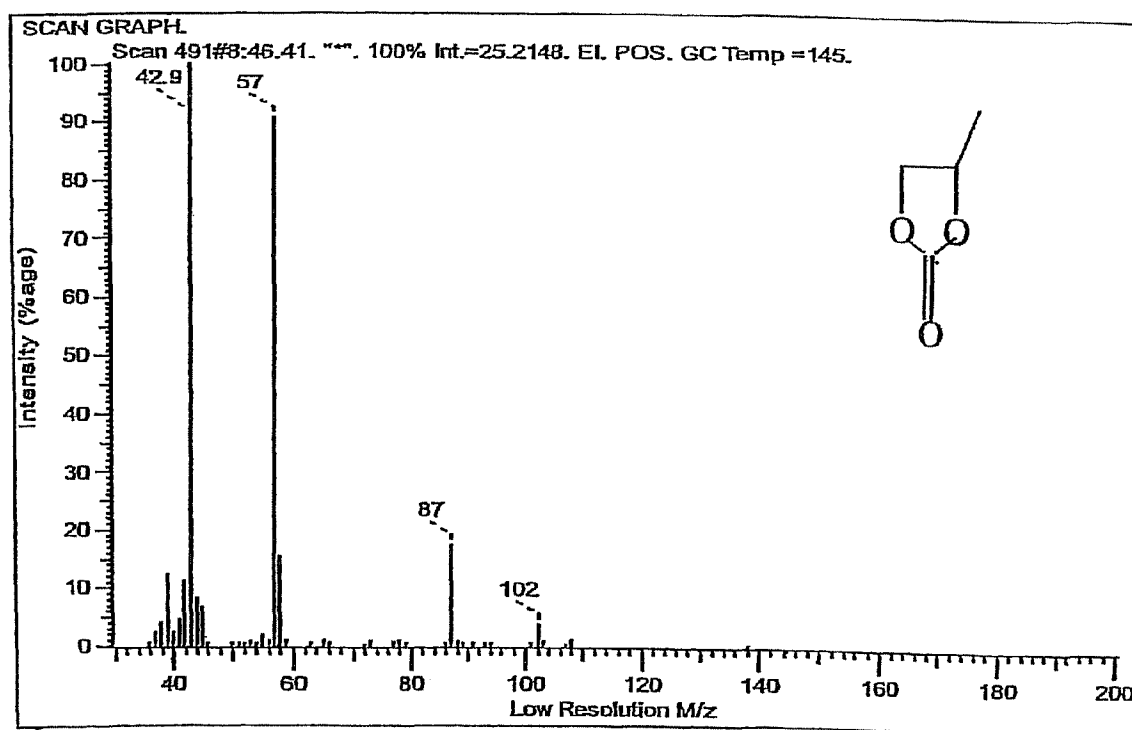


FIG. 15

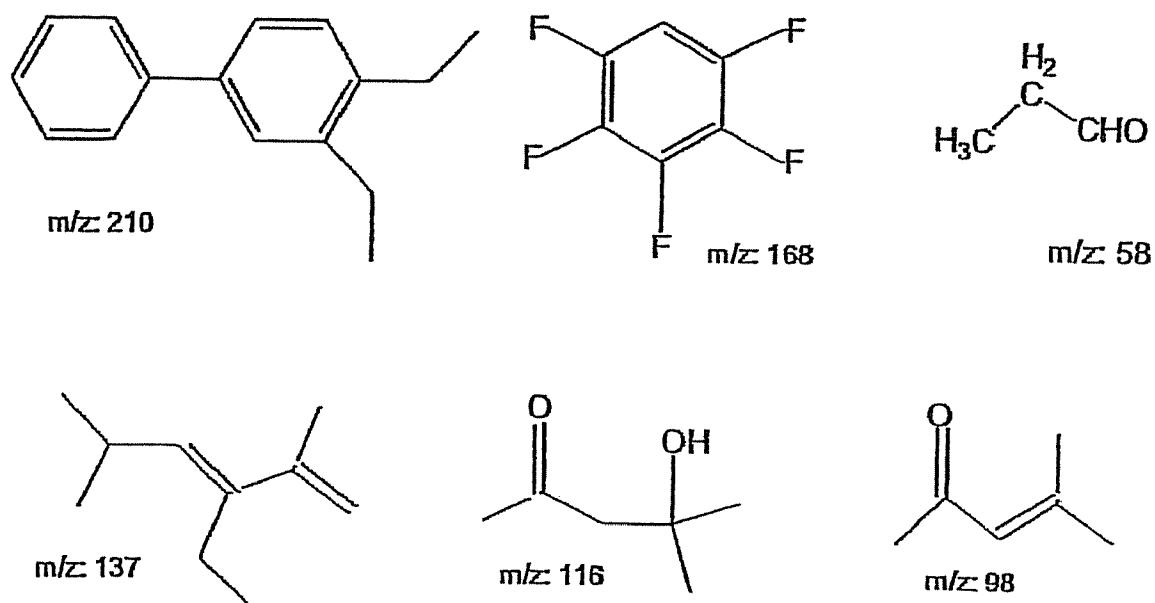
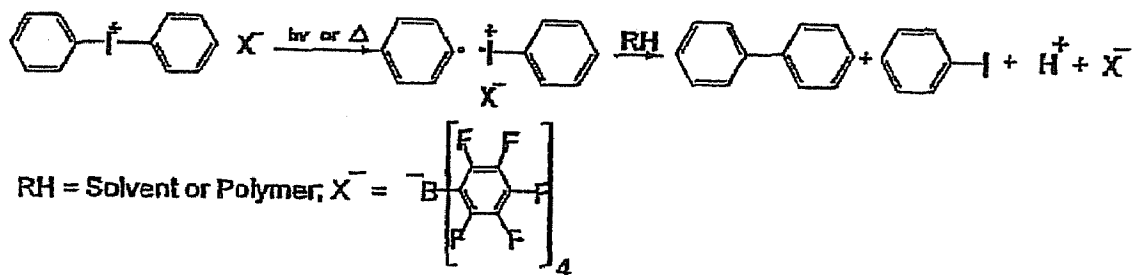


FIG. 16

## (a) Decomposition of the PAG



## (b) Acid catalyzed decomposition of PPC

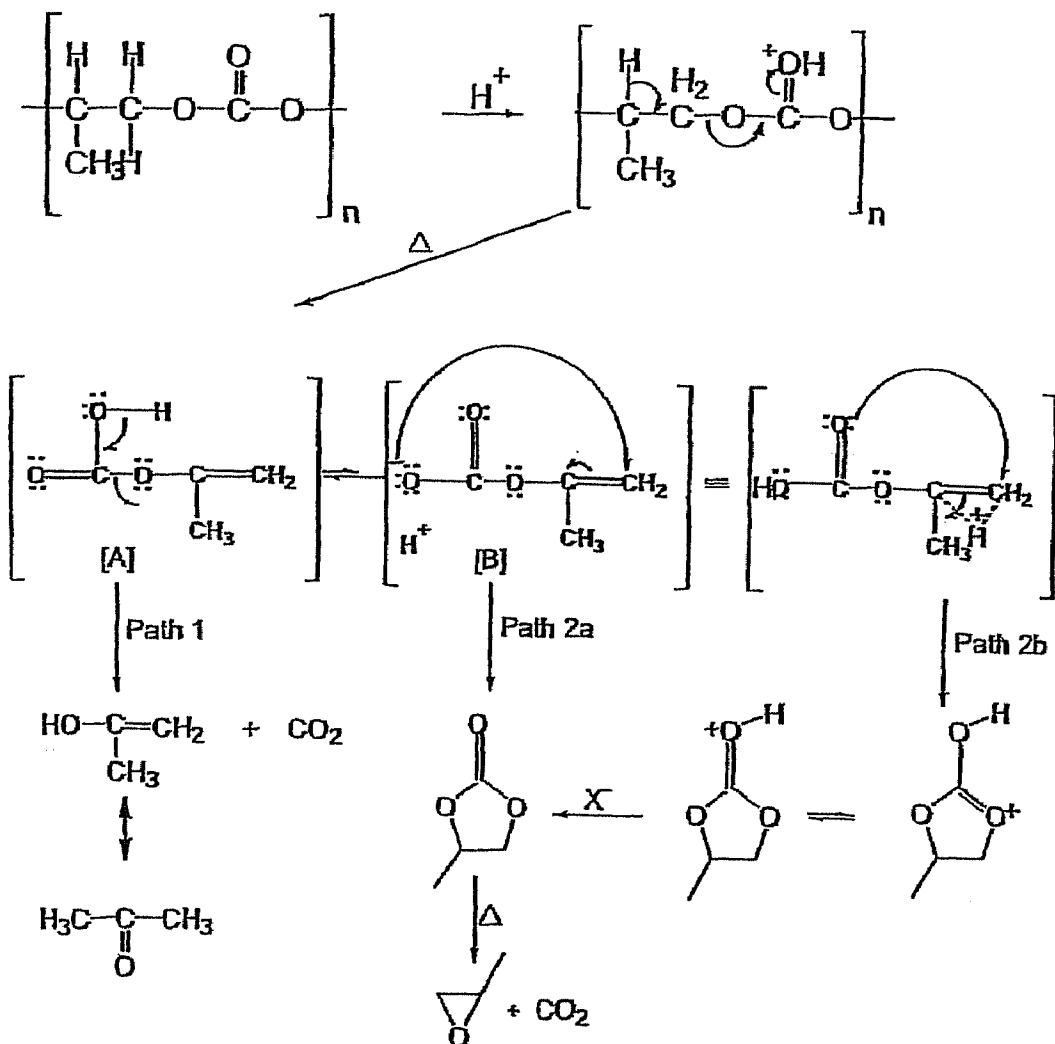


FIG. 17

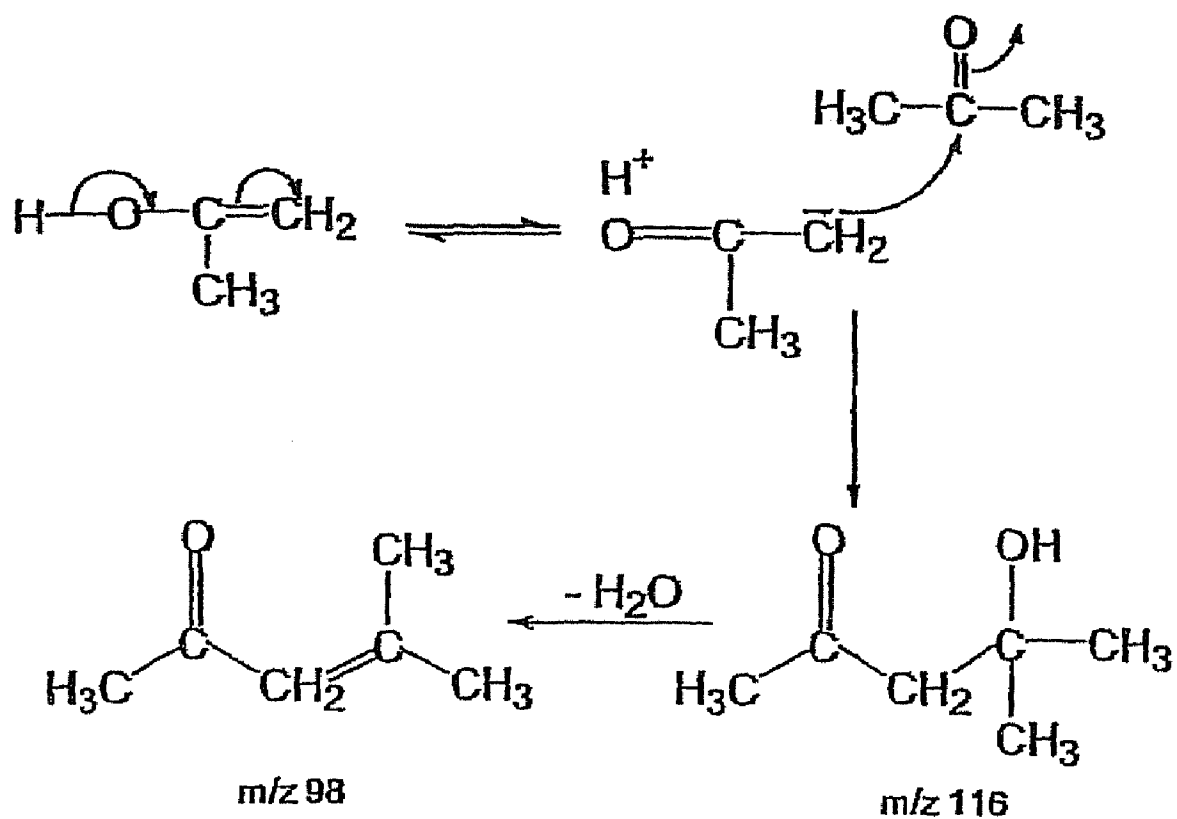


FIG. 18

1

# SACRIFICIAL COMPOSITIONS AND METHODS OF FABRICATING A STRUCTURE USING SACRIFICIAL COMPOSITIONS

## CLAIM OF PRIORITY

This application claims priority to and is a divisional of co-pending U.S. patent application No. 11/398,183, filed Apr. 5, 2006, which is a divisional of co-pending U.S. utility patent application entitled "Sacrificial Compositions, Methods of Use Thereof, and Methods of Decomposition Thereof," having application Ser. No. 10/699,330 and filed on Oct. 31, 2003, which claims priority to U.S. provisional application entitled "Novel Selective-Temperature Sacrificial Polymeric Materials" having Application No. 60/423,013, filed on Nov. 1, 2002, all of which are entirely incorporated herein by reference.

## STATEMENT REGARDING FEDERALLY SPONSORED RESEARCH OR DEVELOPMENT

The U.S. government may have a paid-up license in this invention and the right in limited circumstances to require the patent owner to license others on reasonable terms as provided for by the terms of MDA awarded by the National Science Foundation (Grant #DMI-9980804) of the U.S. Government.

## TECHNICAL FIELD

The present invention is generally related to sacrificial compositions, and, more particularly, is related to sacrificial polymers and their decomposition catalysts, methods of use thereof, and methods of decomposition of the sacrificial polymers.

## BACKGROUND

A wide spectrum of microelectronic and microelectromechanical systems (MEMS) applications has increased the need for lower-temperature, thermally decomposable sacrificial materials. This includes fabrication of air-gaps in electrical interconnects, MEMS, microfluidic devices, and micro-reactors.

The formation of air-gaps is important in electrical interconnects because it lowers the effective dielectric constant of the matrix. The fabrication of buried air channels is useful for the creation of vias in multi-level wiring boards, micro-display boards with high resolution, and ink-jet printer heads. In MEMS technology, the fabrication of micro-air cavities may alleviate the stress associated with thermal expansion of materials and also can act as a temperature-activated release material.

Microfluidic devices and microreactors, fabricated with air-gap technology can be used for miniature-scale chemical syntheses, medical diagnostics, and micro-chemical analysis and sensors. In such devices, liquids and gases are manipulated in microchannels with cross-sectional dimensions on the order of tens to hundreds of micrometers. Processing in such microchannel devices offers a number of advantages including low reagent and analyte consumption, highly compact and portable systems, fast processing times, and the potential for disposable systems.

In spite of all of their promise, however, microfluidic devices are currently being used in a limited number of applications and are in general still rather simple devices in terms of their operational complexity and capabilities. For example,

2

in terms of making truly portable microanalytical systems, one of the current difficulties involves the simple integration of electronic (e.g., sensing methods) and fluidic elements into the same device. One of the most important issues, controlling this ability to integrate functions into the same device, and thus controlling the level of functionality of a microfluidic device, is the method used to fabricate the structure.

The applications for a microfluidic device require the formation of buried microchannels in several different materials at a variety of temperatures. Polycarbonates have been used as a sacrificial material in fabricating nanofluidic devices by electron beam lithography. C. K. Harnett, et al., *J. Vac. Sci. Technol. B.*, vol. 19(6), p. 2842, 2001. Air-gaps have been also fabricated using the hot-filament chemical vapor deposition of polyoxymethylene as a sacrificial layer. L. S. Lee, et al., *Electrochem. and Solid State Lett.*, vol. 4, p. G81, 2001. Further, highly structured, dendritic material, specifically hyperbranched polymers, have been used as a dry-release sacrificial material in the fabrication of a cantilever beam. H.-J. Suh, et al., *J. Microelectromech. Syst.*, Vol. 9(2), pp. 198-205, 2000. Previous work has also fabricated air-gaps using non-photosensitive sacrificial polymers that decompose in the range 250-425° C. P. A. Kohl, et al., *Electrochemical and Solid State Lett.*, vol.1, p.49, 1998; D. Bhusari, et al., *J. Micromech. Microeng.*, vol.10(3), p. 400, 2001.

FIGS. 1A-1H are cross-sectional views that illustrate a previously proposed method **100** for forming a buried air cavity using a non-photosensitive sacrificial material. FIG. 1A illustrates a substrate **10**, prior to having a non-photosensitive sacrificial material **12** disposed thereon by, for example, spin-coating as in FIG. 1B. FIG. 1C illustrates a hard mask **14** disposed on the sacrificial material **12**. FIG. 1D illustrates a photolithographed and etched mask portion **16** disposed on the hard mask **14**, while FIG. 1E illustrates the removal of the mask portion **16** and portions of the sacrificial polymer material **12** exposed to the plasma etching. The hard mask **14** is then removed, as shown in FIG. 1F.

FIG. 1G illustrates the formation of the overcoat layer **18** onto the sacrificial polymer **12** and the substrate **10**. FIG. 1H illustrates the decomposition of the sacrificial polymer **12** to form air-regions **20**. The sacrificial polymer **12** has conventionally been decomposed by heating the sacrificial polymer **12** to a temperature sufficient to decompose the polymer (e.g., about 300-425° C.). Prior sacrificial polymers have required high decomposition temperatures (e.g., about 300-425° C.), which limits the types of overcoat materials and substrates that may be used, as the overcoat and substrate materials must be able to withstand the high temperatures needed to decompose the sacrificial polymers. Even the use of a polycarbonate polymer as the sacrificial polymer lowers the decomposition temperature to about 250-280° C.

Thus, a heretofore unaddressed need exists in the industry to address the aforementioned deficiencies and inadequacies.

## SUMMARY OF THE INVENTION

Briefly described, embodiments of this disclosure, among others, include polymer compositions and methods of use thereof, particularly, methods of fabricating a structure using sacrificial compositions. An exemplary composition, among others, includes a polymer and a catalytic amount of a negative tone photoinitiator.

Methods of fabricating a structure are also provided. One exemplary method, among others, includes the steps of: disposing a composition onto a surface, the composition including a sacrificial polymer and a photoacid generator; exposing at least a portion of the composition of the composition to



energy; and removing a portion of the composition to form an air-gap in the composition, the removed portion corresponding to the portions exposed to the energy. Another exemplary method of fabricating a structure includes following steps: disposing a composition onto a surface, the composition including a sacrificial polymer and a catalytic amount of a photoacid generator; exposing a portion of the composition to energy; and removing the portion of the composition exposed to energy to form an air-gap in the composition via heating the composition to about 100 to 180° C.

Other compositions, methods, features, and advantages will be, or become, apparent to one with skill in the art upon examination of the following drawings and detailed description. It is intended that all such additional compositions, methods, features, and advantages be included within this description, be within the scope of the present invention, and be protected by the accompanying claims.

### BRIEF DESCRIPTION OF THE DRAWINGS

Many aspects of this disclosure can be better understood with reference to the following drawings. The components in the drawings are not necessarily to scale, emphasis instead being placed upon clearly illustrating the principles of this disclosure. Moreover, in the drawings, like reference numerals designate corresponding parts throughout the several views.

FIGS. 1A through 1H are cross-sectional views that illustrate a previously proposed method for forming buried air-gaps formation using a non-photosensitive sacrificial material.

FIGS. 2A through 2F are cross-sectional views that illustrate a representative method for forming buried air-gaps using a polymer that undergoes acid-catalyzed decomposition.

FIGS. 3A through 3F are cross-sectional views that illustrate an alternative representative method for forming buried air-gaps using a polymer that undergoes acid-catalyzed decomposition.

FIGS. 4A through 4F illustrate cross-sectional views that illustrate an alternative representative method for forming buried air cavities using a polymer that undergoes negative-tone stabilization to make the sacrificial polymer more difficult to decompose.

FIG. 5 illustrates chemical structures of exemplary sacrificial polymers used in the methods of FIGS. 2-4.

FIG. 6 illustrates chemical structures of exemplary photoacid generators used in the methods of FIGS. 2-3.

FIG. 7 illustrates chemical structures of further exemplary photoacid generators used in the methods of FIGS. 2-3.

FIG. 8 is a thermogravimetric plot of various disclosed sacrificial compositions.

FIG. 9 is a scanning electron microscope (SEM) micrograph of the lithographic images of an air-gap formed from a representative disclosed sacrificial composition.

FIG. 10 is a SEM micrograph of an air-gap formed from an alternative sacrificial composition.

FIG. 11 is a SEM micrograph of an air-gap formed from an alternative sacrificial composition.

FIG. 12 is a graph of infrared (IR) spectra for a representative disclosed sacrificial composition during the various stages of a representative method of forming air channels.

FIG. 13 illustrates mass spectrometry (MS) scans of a representative disclosed sacrificial composition after the composition has been exposed to ultraviolet (UV) radiation.

FIG. 14 illustrates MS scans of a representative disclosed sacrificial composition without exposure to UV radiation.

FIG. 15 illustrates selected MS scans for a representative sacrificial polymer, and its related fragmentation products.

FIG. 16 depicts the chemical structures of other representative volatile compounds that evolve from the degradation of a representative sacrificial polymer formulation.

FIG. 17 depicts the chemical structures of a mechanism for the acid-catalyzed decomposition of a representative sacrificial polymer.

FIG. 18 depicts the chemical structures of a mechanism for the aldol condensation products evolved during the acid-catalyzed degradation reaction depicted in FIG. 17.

### DETAILED DESCRIPTION

In general, polymers, methods of use thereof, structures formed therefrom, and methods of decomposition thereof, are disclosed. Embodiments of the polymer can be used to form air-gaps in electrical interconnects, microelectromechanical (MEMS), microfluidic devices, and micro-reactors. In addition, methods are disclosed in which sacrificial compositions are decomposed at reduced temperatures than traditionally used.

Embodiments of the disclosed sacrificial composition include, but are not limited to, a sacrificial polymer and one or more positive tone or negative tone components. The positive tone component can include a photoacid generator, for example.

In general, the photoacid generator can be used to make the sacrificial polymer easier to remove (e.g., less stable towards a solvent). For example, half of a layer of a sacrificial composition (e.g., a sacrificial polymer and a positive tone component) is exposed to energy, either in the form of thermal energy (e.g., increased temperature) or optical energy (e.g., ultraviolet (UV) light, near-ultraviolet light, and/or visible light), while the other half is not exposed. Subsequently, the entire layer is exposed to a solvent or heat and the solvent or heat dissolves the layer exposed to the energy.

Although not intending to be bound by theory, upon exposure to optical energy, a positive tone photoacid generator generates an acid. Then, upon exposure to a base or an increased temperature, the dissolution of the sacrificial polymer is increased relative to sacrificial composition not exposed to the optical or thermal energy. As a result, a mask, for example, can be used to fabricate three-dimensional structures from the sacrificial composition by removing the exposed sacrificial polymer.

In general, negative tone compositions can be used making the sacrificial polymer more difficult to remove (e.g., more stable towards a solvent or heat than normally would dissolve the sacrificial polymer). For example, half of a layer of a sacrificial composition (including a sacrificial polymer and a negative tone photoinitiator) is exposed to optical energy, while the other half is not exposed. Subsequently, the entire layer is exposed to a solvent or heat and the solvent or heat dissolves the layer not exposed to the optical energy.

More specifically, the area exposed includes a cross-linked photodefinable polymer, while portions not exposed include an uncross-linked photodefinable polymer. The uncross-linked photodefinable polymer can be removed with the solvent leaving the cross-linked photodefinable polymer behind (e.g., a photodefinable three-dimensional structure).

Although not intending to be bound by theory, upon exposure to optical energy, one type, among others, of the negative tone photoinitiator can generate free radicals that initiate cross-linking reactions between the sacrificial polymers to form a cross-linked photodefinable polymer. Therefore, a mask, for example, can be used to fabricate photodefinable three-dimensional structures from the photodefinable polymer by removing the uncross-linked photodefinable polymer.

In general, the sacrificial composition can be used in areas such as, but not limited to, microelectronics (e.g., microprocessor chips, communication chips, and optoelectronic chips),

microfluidics, sensors, analytical devices (e.g., microchromatography), as a sacrificial material to create three-dimensional structures that can be subsequently have formed therein air-regions (also referred to herein interchangeably as “air-gaps,” “air cavities,” and/or “air channels”) by thermally decomposing the sacrificial polymer. In addition, the sacrificial polymer can be used as an insulator, for example.

For embodiments using the sacrificial composition to create air-regions having three-dimensional structures, the decomposition of the sacrificial composition should produce gas molecules small enough to permeate one or more of the materials surrounding the sacrificial composition (e.g., an overcoat layer). In addition, the sacrificial composition preferably decomposes slowly, so as not to create undue pressure build-up while forming the air-region within the surrounding materials. Furthermore, the sacrificial composition desirably has a decomposition temperature less than the decomposition or degradation temperature of the surrounding material. The sacrificial composition also desirably has a decomposition temperature above the deposition or curing temperature of an overcoat material but less than the degradation temperature of the components in the structure in which the sacrificial composition is being used.

The sacrificial polymer can include compounds such as, but not limited to, polynorbornenes, polycarbonates, functionalized compounds of each, a copolymer of polynorbornene and polynorbornene carbonate, and combinations thereof. The polynorbornene can include, but is not limited to, an alkenyl-substituted norbornene (e.g., cyclo-acrylate norbornene). The polycarbonate can include, but is not limited to, polypropylene carbonate (PPC), polyethylene carbonate (PEC), polycyclohexane carbonate (PCC), polycyclohexanepropylene carbonate (PCPC), and polynorbornene carbonate (PNC), and combinations thereof. The chemical structures of representative sacrificial polymers included in the disclosed sacrificial composition are depicted in FIG. 5. Specific polycarbonates that may be used as the disclosed sacrificial polymer include, for example, poly[(oxycarbonyloxy-1,1,4,4-tetramethylbutane)-alt-oxycarbonyloxy-5-norbornene-2-endo-3-endo-dimethane]; poly[(oxycarbonyloxy-1,4-dimethylbutane)-alt-(oxycarbonyloxy-5-norbornene-2-endo-3-endo-dimethane)]; poly[(oxycarbonyloxy-1,1,4,4-tetramethylbutane)-alt-(oxycarbonyloxy-p-xylene)]; and poly[(oxycarbonyloxy-1,4-dimethylbutane)-alt-(oxycarbonyloxy-p-xylene)]. In general, the molecular weight of the disclosed sacrificial polymers is from about 10,000 to 200,000.

The sacrificial polymer can be from about 1% to 50% by weight of the sacrificial composition. In particular, the sacrificial polymer can be from about 5% to 15% by weight of the sacrificial composition.

As mentioned above, the sacrificial composition can include either a negative tone component and/or a positive tone component. The negative tone component can include compounds that generate a reactant that would cause the crosslinking in the sacrificial polymer. The negative tone component can include compounds, such as, but not limited to, a photosensitive free radical generator. Alternative negative tone components can be used, such as photoacid generators (e.g., in epoxide-functionalized systems).

A negative tone photosensitive free radical generator is a compound which, when exposed to light breaks into two or more compounds, at least one of which is a free radical. In particular, the negative tone photoinitiator can include, but is not limited to, bis(2,4,6-trimethylbenzoyl)-phenylphosphine oxide (Irgacure 819, Ciba Specialty Chemicals Inc.); 2-benzyl-2-dimethylamino-1-(4-morpholinophenyl)-butanone-1 (Irgacure 369, Ciba); 2,2-dimethoxy-1,2-diphenylethane-1-one (Irgacure 651, Ciba); 2-methyl-1[4-(methylthio)-phe-

nyl]-2-morpholinopropan-1-one (Irgacure 907, Ciba); benzoin ethyl ether (BEE, Aldrich); 2-methyl-4'-(methylthio)-2-morpholino-propionophenone; 2,2'-dimethoxy-2-phenyl-acetophenone (Irgacure 1300, Ciba); 2,6-bis(4-azidobenzylidene)-4-ethylcyclohexanone (BAC-E), and combinations thereof.

The positive tone component can include, but is not limited to, photoacid generator(s). More specifically, the positive tone photoacid generator can include, but is not limited to, nucleophilic halogenides (e.g., diphenyliodonium salt, diphenylfluoronium salt) and complex metal halide anions (e.g., triphenylsulphonium salts). In particular, the photoacid generator can be tetrakis(pentafluorophenyl)borate-4-methylphenyl[4-(1-methylethyl)phenyl]iodonium (DPI-TPFPB); tris(4-t-butylphenyl)sulfonium tetrakis(pentafluorophenyl) borate (TTBPS-TPFPB); tris(4-t-butylphenyl)sulfonium hexafluorophosphate (TTBPS-HFP); triphenylsulfonium triflate (TPS-Tf); bis(4-tert-butylphenyl)iodonium triflate (DTBPI-Tf); triazine (TAZ-101); triphenylsulfonium hexafluoroantimonate (TPS-103); Rhodasil™ Photoinitiator 2074 (FABA); triphenylsulfonium bis(perfluoromethanesulfonyl) imide (TPS-N1); di-(p-t-butyl) phenyliodonium bis(perfluoromethanesulfonyl) imide (DTBPI-N1); triphenylsulfonium; tris(perfluoromethanesulfonyl)methide (TPS-C1); di-(p-t-butylphenyl)iodonium tris(perfluoromethanesulfonyl)methide (DTBPI-C1); and combinations thereof, the chemical structures of which are depicted in FIGS. 6 and 7.

The photoacid generator can be from about 0.5% to 5% by weight of the sacrificial composition. In particular, the photoacid generator can be from about 1% to 3% by weight of the sacrificial composition.

The remaining percentage of the sacrificial composition not accounted for in the photoacid generator and sacrificial polymer (e.g., from about 50% to about 99%) can be made up with solvent, such as, but not limited to, mesitylene, N-methyl-2-pyrrolidinone, propylene carbonate, anisole, cyclohexanone, propylene glycol monomethyl ether acetate, N-butyl acetate, diglyme, ethyl 3-ethoxypropionate, and combinations thereof.

Exemplary sacrificial polymers include those chemical structures depicted in FIG. 5. When using sacrificial polymers to fabricate air-gap structures, it is helpful to know the degradation temperatures of the sacrificial polymers. Table 1 below shows the half-decomposition temperatures,  $T_{1/2d}$  (where  $T_{1/2d}$  is the temperature at 50% weight loss) of exemplary sacrificial polymers under various experimental conditions, determined using dynamic thermogravimetric analysis (TGA).

TABLE 1

Half-decomposition temperatures, $T_{1/2d}$ (where $T_{1/2d}$ is the temperature at 50% weight loss) of exemplar sacrificial polymers under various conditions.						
Entry	Polycarbonate	Solvent	PAG	PIAD $T_{1/2d}$ (° C.)	TIAD $T_{1/2d}$ (° C.)	TD $T_{1/2d}$ (° C.)
1.	12% PPC	Anisole	DPI-TPFPB	102 ± 2	180 ± 2	212 ± 2
2.	20% PPC	Anisole	TTBPS-TPFPB	125 ± 2	214 ± 2	212 ± 2
3.	10% PEC	Propylene carbonate	DPI-TPFPB	114 ± 2	189 ± 2	182 ± 2
4.	10% PEC	Propylene carbonate	TTBPS-HFP	117 ± 2	184 ± 2	182 ± 2
5.	12% PCC	Anisole	DPI-TPFPB	177 ± 2	199 ± 2	270 ± 2
6.	35% PCC	Anisole	TTBPS-TPFPB	191 ± 2	252 ± 2	270 ± 2
7.	20% PCPC	Anisole	DPI-TPFPB	160 ± 2	194 ± 2	252 ± 2

TABLE 1-continued

Half-decomposition temperatures, $T_{1/2d}$ (where $T_{1/2d}$ is the temperature at 50% weight loss) of exemplar sacrificial polymers under various conditions.						
Entry	Polycarbonate	Solvent	PAG	PIAD $T_{1/2d}$ (° C.)	TIAD $T_{1/2d}$ (° C.)	TD $T_{1/2d}$ (° C.)
8.	20% PCPC	Anisole	TTBPS-TPFPB	$186 \pm 2$	$249 \pm 2$	$252 \pm 2$
9.	12% PNC	Anisole	DPI-TPFPB	$207 \pm 2$	$194 \pm 2$	$305 \pm 2$
10.	50% PNC	Anisole	TTBPS-TPFPB	$196 \pm 2$	$291 \pm 2$	$304 \pm 2$

Now having described the sacrificial composition in general, the following describes exemplar embodiments for using the sacrificial composition to produce three-dimensional structures, where the three-dimensional structures can be decomposed to form air-regions (e.g., a gas filled region substantially excluding a solid or liquid material or a vacuum-region).

In general, disposing a layer of the sacrificial composition onto a substrate and/or layer of material on the substrate can produce a three-dimensional structure. A mask is disclosed on or above the sacrificial composition or portions thereof that encodes the shape of the three-dimensional structure, as described below. After exposing the sacrificial composition through the mask to optical and/or thermal energy and removing the unexposed sacrificial composition (negative tone) or the exposed sacrificial composition (positive tone), the three-dimensional structure is formed.

The mask encodes a density profile that defines the three-dimensional structure. Upon exposure of the mask to optical and/or thermal energy, a known amount of the energy is allowed to pass through portions of the mask. The design of the mask is used to control the amount of energy allowed to pass through the mask. In particular, the mask can be designed to control the amount of energy allowed to pass through the mask as a function of the position on the mask. Thus, the mask can be designed and used to produce the three-dimensional structure from the sacrificial composition by altering the amount of energy allowed to pass through the mask as a function of the position on the mask. The mask can be formed by methods known in the art (e.g., U.S. Pat. No. 4,622,114).

The three-dimensional structures (and the corresponding air-regions) can have cross-sectional areas section such as, but not limited to, non-rectangular cross-sections, asymmetrical cross-sections, curved cross sections, arcuate cross sections, tapered cross sections, cross sections corresponding to an ellipse or segment thereof, cross sections corresponding to a parabola or segment thereof, cross sections corresponding to a hyperbola or segment thereof, and combinations thereof. For example, the three-dimensional structures can include, but are not limited to, non-rectangular structures, non-square structures, curved structures, tapered structures, structures corresponding to an ellipse or segment thereof, structures corresponding to a parabola or segment thereof, structures corresponding to a hyperbola or segment thereof, and combinations thereof. In addition, the three-dimensional structures can have cross-sectional areas having a spatially-varying height. Although not illustrated, a non-rectangular, tapered, and asymmetrical air-region can be formed in conjunction with other air-regions and/or air-channels to form microfluidic devices, sensors, and analytical devices, for example.

An exemplar embodiment of a device with air-regions **250** formed therein is shown in cross-section in FIG. **2F**. The device may include a substrate **210**, and an overcoat layer **240**.

A substrate **210** on which the sacrificial composition is disposed can be used in systems such as, but not limited to, microprocessor chips, microfluidic devices, sensors, analytical devices, and combinations thereof. Thus, the substrate **210** can be made of materials appropriate for the particular desired system or device. Exemplar materials, however, include, but are not limited to, glasses, silicon, silicon compounds, germanium, germanium compounds, gallium, gallium compounds, indium, indium compounds, or other semiconductor materials and/or compounds. In addition, the substrate **210** can include non-semiconductor substrate materials, including any dielectric material, metals (e.g., copper and aluminum), ceramics, or organic materials found in printed wiring boards, for example.

The overcoat layer **240** can be a modular polymer that includes the characteristic of being permeable or semi-permeable to the decomposition gases produced by the decomposition of a sacrificial polymer while forming the air-region **250**. In addition, the overcoat layer **240** has elastic properties so as to not rupture or collapse under fabrication and use conditions. Further, the overcoat layer **240** is stable in the temperature range in which the sacrificial composition decomposes.

Examples of the overcoat layer **240** include compounds such as, but not limited to, polyimides, polynorbornenes, epoxides, polyarylenes ethers, polyarylenes, inorganic glasses, and combinations thereof. More specifically the overcoat layer **240** includes compounds such as Amoco Ultradel™ 7501, Promerus Avatrel™ Dielectric Polymer, DuPont 2611, DuPont 2734, DuPont 2771, DuPont 2555, silicon dioxide, silicon nitride, and aluminum oxide. The overcoat layer **240** can be deposited onto the substrate **210** using techniques such as, for example, spin coating, doctor-blading, sputtering, lamination, screen or stencil-printing, chemical vapor deposition (CVD), metalorganic chemical vapor deposition (MOCVD), and plasma-based deposition systems.

It should be noted that additional components could be disposed on and/or within the substrate **210**, the overcoat layer **240**, and/or the air-region **250**. In addition, the additional components can be included in any structure having air-regions as described herein. The additional components can include, but are not limited to, electronic elements (e.g., switches and sensors), mechanical elements (e.g., gears and motors), electromechanical elements (e.g., movable beams and mirrors), optical elements (e.g., lens, gratings, and mirror), opto-electronic elements, fluidic elements (e.g., chromatograph and channels that can supply a coolant), and combinations thereof.

FIGS. **2A-2F** are cross-sectional views that illustrate a representative method **200** for forming an air-region using a sacrificial composition, the sacrificial composition including a sacrificial polymer and a positive tone component. FIG. **2A** illustrates a substrate **210**, prior to having a sacrificial material **220** disposed thereon by, for example, spin-coating as shown in FIG. **2B**. FIG. **2C** illustrates a mask **230** disposed on or above the sacrificial material **220** and the sacrificial material **220** being subjected to energy/radiation. While FIG. **2C** depicts the energy being applied in the form of UV radiation, other forms of optical energy or thermal energy can be applied.

FIG. **2D** illustrates the removal of selected portions of the sacrificial composition **220** that was exposed to the UV radiation.

tion of FIG. 2C through the mask 230. The mask 230 encodes an optical density profile that defines the cross-section of the air-region 250. The removed portions of the sacrificial composition 220 are removed by heating the sacrificial composition. The temperature used to remove irradiated portions of the sacrificial material 220 is lower than the methods of the prior art, due to the presence of a photoacid generator, described above, in the sacrificial composition 220. The temperature required to substantially decompose the irradiated sacrificial polymer 220 in the step shown in FIG. 2C is about 50° C. to 400° C. For example, the temperature may be about 110° C.

FIG. 2E illustrates the formation of the overcoat layer 240 onto the sacrificial polymer 220 and exposed substrate 210. FIG. 2F illustrates the decomposition of the sacrificial polymer 220 to form the air-regions 250. The temperature required to substantially decompose the sacrificial polymer in the step shown in FIG. 2F is about 100° C. to 250° C. For example, the temperature may be about 170° C.

FIGS. 3A-3F are cross-sectional views that illustrate an alternative method 300 for forming an air-region 250 using a sacrificial composition, the sacrificial composition including a sacrificial polymer and a positive tone component. FIG. 3A illustrates a substrate 210, prior to having a sacrificial material 220 disposed thereon by, for example, spin-coating as shown in FIG. 3B. FIG. 3C illustrates a mask 230 disposed on or above the sacrificial material 220 and the sacrificial material 220 being subjected to energy/radiation. While FIG. 3C depicts the energy being applied in the form of UV radiation, other forms of optical energy, or thermal energy, can be applied.

FIG. 3D illustrates the cross-linked sacrificial composition region 225 after exposure through the mask 230 to optical energy. FIG. 3E illustrates the formation of the overcoat layer 240 onto the sacrificial polymer 220 and cross-linked composition regions 225. FIG. 3F illustrates the decomposition of the un-crosslinked sacrificial polymer 220 to form the air-regions 250. The removed portions of the sacrificial composition 220 are removed by heating the sacrificial composition. The temperature used to remove irradiated portions of the sacrificial material 220 in FIG. 3F is lower than the methods of the prior art, due to the presence of a photoacid generator, described above, in the sacrificial composition 220.

FIGS. 4A through 4F are cross-sectional views that illustrate a representative process 400 for fabricating the air-regions 250 illustrated in FIG. 4F. It should be noted that for clarity, some portions of the fabrication process are not included in FIGS. 4A through 4F. As such, the following fabrication process is not intended to be an exhaustive list that includes all steps required for fabricating the air-regions 250. In addition, the fabrication process is flexible because the process steps may be performed in a different order than the order illustrated in FIGS. 4A through 4F, or some steps may be performed simultaneously.

FIG. 4A illustrates the substrate 210. FIG. 4B illustrates the substrate 210 having the sacrificial composition 260 (negative tone) disposed thereon. The sacrificial composition 260 can be deposited onto the substrate 210 using techniques such as, for example, spin coating, doctor-blading, sputtering, lamination, screen or stencil-printing, melt dispensing, chemical vapor deposition (CVD), metalorganic chemical vapor deposition (MOCVD), and plasma-based deposition systems.

FIG. 4C illustrates a mask 230 disposed on or above the sacrificial composition 260. The mask 230 encodes an optical density profile that defines the cross-section of the photodefinable air-region 250.

FIG. 4D illustrates the cross-linked sacrificial composition region 270 after exposure through the mask 230 to optical energy, after the removal of the selected uncross-linked sacrificial composition material 270. The uncross-linked sacrificial composition material 260 can be removed by dissolution in a liquid, such as a solvent and/or heat, for example, or by any other method that can remove or dissolve the polymer component of the sacrificial composition material 260.

FIG. 4E illustrates the formation of the overcoat layer 240 onto the cross-linked sacrificial composition region 270. The overcoat layer 240 can be deposited onto the substrate 210 using techniques such as, for example, spin coating, doctor-blading, sputtering, lamination, screen or stencil-printing, melt dispensing, chemical vapor deposition (CVD), metalorganic chemical vapor deposition (MOCVD), and plasma-based deposition systems.

FIG. 4F illustrates the decomposition of the cross-linked sacrificial composition region 270 to form the air-region 250. The cross-linked sacrificial composition region 270 can be decomposed by heating the cross-linked sacrificial composition 270 to a temperature sufficient to decompose the polymer (e.g., about 250° C.).

Thermogravimetric Analysis of Representative Sacrificial Composition Film

FIG. 8 illustrates the dynamic thermogravimetric (TGA) results for an exemplar sacrificial composition of a sacrificial polymer, polypropylene carbonate (PPC), and a photoacid generator (PAG). The composition was comprised of approximately 12 wt % PPC and 5 wt % DPI-TPFPB of PPC ("PPC:PAG"). The representative samples were prepared by spin-coating the polymer in anisole onto a silicon wafer, soft baking on a hotplate at about 110° C. for 10 minutes to evaporate the solvent and then exposing, where applicable, to 1 J/cm<sup>2</sup> of 240 nm UV irradiation. The samples were then removed from the silicon and analyzed by dynamic TGA ramping from 30 to 450° C. at the rate of 1° C. under a nitrogen atmosphere.

FIG. 8 shows the TGA thermograms for (a) PPC:PAG after UV irradiation, (b) PPC:PAG without UV irradiation, (c) PPC without PAG, and (d) PPC without PAG but with UV irradiation. As shown in FIG. 8(a), the onset temperature for the photo-acid induced decomposition was found to be 80° C., and the  $T_{1/2,d}$  was 100° C. The decomposition was complete at 230° C. with less than 3 wt % residue in the TGA pan.

When the sacrificial polymer material was not UV irradiated (FIG. 8(b)), decomposition occurred via the thermolytically induced PAG acid generation once the temperature reached the decomposition temperature of the PAG. The decomposition occurred in a narrow temperature range from the onset at 168° C. At 230° C., the thermal-acid induced decomposition was complete with less than 1.5 wt % residue.

The decomposition behavior of PPC without PAG is shown in FIG. 8(c), and the  $T_{1/2,d}$  was found to be 210° C. At 230° C., 19 wt % of the mass remained in the pan. The decomposition was complete at 287° C. with 3 wt % residue. At 350° C., the weight percentages of the residues were 1.79%, 0.12% and 0.37% for (a), (b), and (c), respectively, under the same experimental conditions.

The decomposition behavior of the sacrificial polymer (here, PPC) film without PAG but exposed to UV irradiation (FIG. 8(d)) was found to be similar to that of the non-UV irradiated sacrificial polymer film. This shows that no change occurred in the sacrificial polymer upon UV irradiation without PAG. Thus, the photo-acid induced decomposition of the polycarbonates significantly lowers the decomposition temperatures; however, it leaves more residue than the thermolytically induced acid-catalyzed decomposition or ther-

molitic decomposition of polypropylene carbonate alone. The residue level can be brought to less than 1 wt % by lowering the percentage of PAG level in the formulation for the acid-catalyzed degradation.

#### Photopatterning and Fabrication of Air-Gaps Using Photosensitive Sacrificial Compositions

Positive-tone patterns obtained using PPC (12 wt %):DPI-TPFPB (5 wt % of PPC) formulation in anisole. The solution was spin-coated onto a silicon and soft-baked on a hotplate at 110° C. for 10 min to achieve a thickness of 1.58  $\mu\text{m}$ . The film was UV irradiated through a clear-field quartz mask with a dose of 1 J/cm<sup>2</sup> (240 nm). The film was dry-developed by post-exposure baking at 110° C. on a hotplate for 3 min. At 110° C., the UV generated acid induced the decomposition of the PPC. FIG. 9 shows scanning electron microscope (SEM) micrographs of lithographic images produced using a dark-field mask. The process was repeated as above except a PPC (20 wt %):DPI-TPFPB (5 wt % of PPC) in anisole formulation was used to achieve a 5.45  $\mu\text{m}$  thick film. The circles and squares are the areas in which the PPC has been decomposed and removed.

#### Examples of Implementations of Disclosed Methods

Air-gaps were fabricated via the process flow using method 200 depicted in FIG. 2 using photosensitive polycarbonates as the sacrificial polymeric material. Two different encapsulating materials were used: Avatrel EPM™ and Avatrel 2000P™ (Promerus, LLC.). First, the photosensitive PPC (12 wt %):DPI-TPFPB (5 wt % of PPC) formulation was spin-coated onto a silicon wafer and soft-baked on a hotplate at 100° C. for 10 minutes to achieve a thickness of 5.45  $\mu\text{m}$ . The photosensitive film was photopatterned using a clear-field mask having lines/space pattern with 70- $\mu\text{m}$  wide lines and 35- $\mu\text{m}$  wide spaces. Post-exposure baking at 110° C. for 1-10 min. decomposed the UV irradiated area. The AVATREL EPM encapsulating material was spin-coated and soft-baked on a hotplate at 80° C. for 5 minutes. The unexposed area was decomposed in a Lindberg horizontal tube furnace at 150° C. for 4 hours under nitrogen. FIG. 10 shows a cross-sectional SEM micrograph image of the 70- $\mu\text{m}$  wide resulting air channel structure encapsulated by 3.9  $\mu\text{m}$  AVATREL EPM™ dielectric material. A similar fabrication process sequence was followed, using AVATREL 2000P™ as the encapsulating material. However, after spin-coating AVATREL 2000P, the film was UV irradiated with 1 J/cm<sup>2</sup> at 365 nm and post-exposure baked in an oven at 110° C. for 30 minutes. Upon baking, the AVATREL 2000P™ polymer was cross-linked. The resulting thickness of the overcoat was 9.3  $\mu\text{m}$ . The unexposed sacrificial polymer was then decomposed in a Lindberg horizontal tube furnace at 170° C. for 1 hour under nitrogen.

FIG. 11 shows an SEM image of the resulting 70- $\mu\text{m}$  wide buried air-channel in AVATREL 2000P™ encapsulant. The air channels were clean from visible debris, which shows the permeability of the volatiles through the encapsulating polymeric material, or "overcoat" material.

Air channels were also fabricated via the process flow using method 300 described in FIG. 3 using a PPC:TTBPS-TPFPB formulation as the sacrificial composition. The PAG, TTBPS-TPFPB, was chosen because it thermally decomposes at a higher temperature than the DPI-TPFPB. A higher decomposition temperature PAG leaves intact the unexposed region of the sacrificial material, while the exposed area is selectively decomposed. The decomposition temperature of TTBPS-TPFPB was found to be 190° C. from the differential scanning calorimetry (DSC). Buried air-channels in AVATREL EPM encapsulant were fabricated using the process sequence described as method 300 of FIG. 3. A PPC (20

wt %):TTBPS-TPFPB (5 wt % of PPC) formulation was spin-coated onto a silicon wafer and soft-baked on a hotplate at 110° C., resulting in a thickness of 5.45  $\mu\text{m}$ . The photosensitive PPC film was irradiated (1 J/cm<sup>2</sup>; 240 nm) through a clear-field mask having with a line/space pattern of 100- $\mu\text{m}$  wide lines and 240- $\mu\text{m}$  wide spaces. The AVATREL EPM encapsulant was then spin-coated and soft-baked at 80° C. for 5 minutes to remove any solvent from the encapsulating layer. The exposed area was then selectively decomposed by heating at 110° C. for 30 minutes in a tube furnace under nitrogen to form air-gaps.

#### FT-IR Analysis of the Decomposition of Representative Sacrificial Composition Film

The acid catalyzed decomposition of PPC (12 wt %):DPI-TPFPB PAG (5 wt % of PPC) was studied by FT-IR. The sacrificial composition was spun onto a NaCl plate and soft-baked on a hotplate at 110° C. for 10 min. The thickness of the film was measured to be 1.45  $\mu\text{m}$ . The Fourier Transform-Infrared (FT-IR) spectrum of the unexposed film was recorded as shown in FIG. 12(a). The film was then exposed to UV light (1 J/cm<sup>2</sup>; 240 nm) and scanned again (FIG. 12(b)). The film was post-exposure baked on a hotplate at 110° C. for 2 h (FIG. 12(c)). In FIG. 12(a), the absorptions at 2990 cm<sup>-1</sup>, 1470 cm<sup>-1</sup> and 1250 cm<sup>-1</sup> were assigned to C—H stretches, C—H bending and C13 C stretching of the PPC respectively. A strong absorption band at 1750 cm<sup>-1</sup> corresponds to C=O stretch of the PPC. On examining the UV-exposed PPC film (FIG. 12(b)), no specific transformation in the chemical structure has occurred, except that the photosensitive acid generator was activated during irradiation. The intensities of all peaks were nearly zero after the final bake at 110° C. (FIG. 12(c)). This is believed to be due to the decomposition of the sacrificial polymer PPC into volatile products.

#### Decomposition of Representative Sacrificial Composition Monitored by Mass Spectrometry

Mass spectrometry (MS) using electron impact ionization was employed to detect the evolved species during the depolymerization and volatilization process for an exemplar sacrificial composition. Three representative samples were analyzed to determine the nature of the chemical species that were produced during degradation. In the first case, the PPC (12 wt %): DPI-TPFPB (5 wt % of PPC) formulation was spin-coated onto a silicon wafer, soft-baked on a hotplate at 110° C. for 10 minutes, then irradiated with 1 J/cm<sup>2</sup> UV irradiation at 240 nm. In the second case, the sample was prepared as above, but was not UV irradiated. In the third case, to study the volatiles evolved from the sacrificial polymer (here, PPC) without a photoacid generator, a sample was made from the solution that contained only PPC. This PPC film was not UV irradiated. The films were removed from the silicon and analyzed by gas chromatograph-mass spectrometer (GC-MS) to study the evolution of volatiles under different conditions. The samples were ramped at 5° C./min from 30 to 110° C. and held for 30 minutes, followed by a ramp at 15° C./min to 300° C. and held for 20 min.

The mass spectrum data for the decomposition of PPC: DPI-TPFPB formulation with UV exposure shows the ion current for a mass-to-charge ratio (m/z) of 58 and 102 which correspond to acetone and propylene carbonate, respectively. Scans of the UV irradiated photosensitive polycarbonate sample are shown in FIG. 13(a), (b) & (c). These scans at different intervals of time during the degradation process reveal the evolution of these two prominent species. FIG. 13(a) also shows other species at m/z 43.1, 40.1 and 36.1 which corresponds to the fragmentation products of acetone. The m/z 58 peak may also correspond to the formation of CH<sub>3</sub>CH<sub>2</sub>CHO. FIG. 13(b) shows other high intensity vola-

## 13

files corresponding to the molecular weights 43, 57, and 87, which are the fragmentation products of the propylene carbonate and may be assigned to ethylene oxide, propylene oxide, and ethylene carbonate, respectively. The mass spectrum at 63° C. is presented in FIG. 13(c). The high intensity peak at 168 m/z can be assigned to a decomposition fragment of the PAG. This matches the mass of pentafluorobenzene (HC<sub>6</sub>F<sub>5</sub>) which is the derivative fragment of the anion of the photoacid, tetrakis(pentafluorophenyl)borate [B(C<sub>6</sub>F<sub>5</sub>)<sub>4</sub>]. This was confirmed by pyrolysing the photoacid (dissolved in anisole) which exhibited a strong mass peak at 168 m/z at 67° C. Tetrakis(pentafluorophenyl)borate anion can decompose on heating to yield pentafluorobenzene and tri(pentafluorophenyl)borate. The reaction includes the volatiles of particular interest that evolved from the degradation of the polycarbonate formulation.

The suspected reaction products are shown in FIG. 16. The biphenyl derivative (m/z 210) and the pentafluorobenzene (m/z 168) resulted from the decomposition of the PAG, DPI-TPFPB. The other fragments having m/z values 137, 116, and 98 are the degradation products of the PPC. At higher temperatures, small amounts of higher molecular weight fragments were observed. These may have resulted from more complex fragmentation patterns.

The mass spectrum of the PPC formulation without UV exposure is shown in FIG. 14 and shows similar fragmentation patterns with that of PPC formulation with UV irradiation. Specifically, m/z peaks at 58 and 102 were observed, corresponding to acetone and propylene carbonate, respectively. The peak at m/z=58 may also correspond to CH<sub>3</sub>CH<sub>2</sub>CHO. This confirms that the acid catalyzed PPC decomposition is the same for photolytic or thermal activation of the PAG. The mass spectrum of the PPC without acid (and without UV irradiation) is shown in FIG. 15. The mass peak at m/z=102 at 145° C. can be assigned to propylene carbonate (as with the acid-catalyzed PPC). The other mass peaks at 87, 57, and 43 can be assigned to the propylene carbonate fragments: ethylene carbonate, propylene oxide, and ethylene oxide, respectively.

#### Mechanism for the Acid-Catalyzed Decomposition of Representative Sacrificial Polymer

In one study, the MS data shows that the degradation of an exemplar sacrificial polymer, e.g., polypropylene carbonate, is initiated by the in-situ acid generated from the PAG, either photolytically or by thermal heating. Without being bound by any theory, it is believed that the degradation of PPC may take place in two stages: the scission reaction of the polymer chain followed by the unzipping reaction. Under pyrolysis at 180° C., it has been observed that polypropylene carbonate yielded propylene carbonate. Further, the propylene carbonate was fragmented as CO<sub>2</sub> and propylene oxide. The evolution of the two prominent species, acetone and propylene carbonate, from the degradation of polypropylene carbonate formulations (with and without UV irradiation) led to the following proposal of a dual-pathway degradation mechanism.

FIG. 17(a) describes the decomposition of the DPI-TPFPB PAG. On exposure to 240 nm UV irradiation, the cation part of the PAG decomposes to produce a proton that pairs with the complex anion to form a protic acid. In this mechanistic scheme, —RH is a donor of protons either from the solvent or the polymer itself. The decomposition of iodonium salts can be achieved by either UV exposure or by thermal breakdown of the cation/anion pair of the iodonium salt. Thus, the iodonium salt can act as a photo-acid generator as well as a latent thermal-acid generator. It has also been reported by that identical reaction products were obtained when onium salts were decomposed either photolytically or thermally. J. V. Crivello, et al., *J. Polym. Sci. Part A: Polym. Chem.*, vol. 21, p. 97, 1983.

## 14

FIG. 17(b) describes the proposed mechanism for the acid-catalyzed decomposition of polypropylene carbonate. During post-bake, the H<sup>+</sup> from the generated acid (H<sup>+</sup>X<sup>-</sup>) protonates the carbonyl oxygen and further rearranges the polar transition state leading to the formation of unstable tautomeric intermediates, [A] and [B]. From the mass spectrum we observed a strong mass peaks at 58 m/z and 102 m/z which was assigned for acetone and propylene carbonate, respectively. The acetone may be formed as the intermediate [A] (Path 1) rearranges and fragmenting as acetone and CO<sub>2</sub>. The formation of propylene carbonate may be attributed to the intramolecular attack of the anion of the intermediate [B] (Path 2a) leading to the formation of the cyclic propylene carbonate. This further breaks down thermally into propylene oxide and CO<sub>2</sub>. The mass peaks at 43 m/z and 57 m/z confirm the formation of ethylene oxide and propylene oxide, respectively.

It is also reasonable to propose an alternative route for the formation of propylene carbonate in this highly acidic environment. The H<sup>+</sup> may activate the terminal double bond (Path 2b), leading to the formation of a cyclic transition state which on further rearrangement and abstraction of a proton by the counter ion X<sup>-</sup>, may yield propylene carbonate. Further confirmation of the degradation via Path 1 is the formation of aldol condensation products. The mass peaks at 116 m/z and 98 m/z may be due to products that were produced due to the aldol condensation (FIG. 18).

## CONCLUSIONS

The fabrication of air-gaps via acid-catalyzed degradation of a sacrificial polymer and a catalytic amount of a photoacid generator has been demonstrated. Based on the FT-IR and mass spectral studies, a detailed mechanism for the polypropylene carbonate and the DPI-TPFPB system has been proposed. The decomposition behavior of several different sacrificial compositions has also been studied using TGA. The low decomposition temperatures provide a mechanism to tailor the decomposition temperature through PAG concentration or sacrificial polymer structure. By providing a processing method to selectively decompose sacrificial composition areas through photoexposure, these materials are a promising candidate for various microelectronic, microfluidic, and MEMS fabrication applications. In particular, the insolubility of the disclosed sacrificial polymers in the solvents used for dielectric materials makes the combination especially valuable.

It should be emphasized that the above-described embodiments of this disclosure are merely possible examples of implementations, and are set forth for a clear understanding of the principles of this disclosure. Many variations and modifications may be made to the above-described embodiments of this disclosure without departing substantially from the spirit and principles of this disclosure. All such modifications and variations are intended to be included herein within the scope of this disclosure and protected by the following claims.

The invention claimed is:

1. A method for fabricating a structure, comprising: disposing a composition onto a surface, wherein the composition includes a sacrificial polymer and a catalytic amount of photoacid generator; exposing a portion of the composition to energy; and removing the portion of the composition exposed to energy to form an air-gap in the composition via heating the composition to about 100 to 180° C., where said sacrificial polymer is chosen from a polycarbonate.

## 15

2. The method of claim 1, further comprising disposing a mask on or above the composition, the mask encoding a profile defining an air-gap to be formed in the composition.

3. The method of claim 1, further comprising removing a portion of the composition not exposed to energy by heating to a temperature of about 175 to 200° C.

4. The method of claim 1, wherein exposing the composition to energy comprises decomposing an organic cation of the photoacid generator, thus generating a strong Brønsted acid.

5. The method of claim 4, wherein exposing the composition to energy comprises thermolytically decomposing the sacrificial polymer with the Brønsted acid.

## 16

6. The method of claim 1, wherein said sacrificial polymer is selected from polypropylene carbonate, polyethylene carbonate, polycyclohexane carbonate, polycyclohexanepropylene carbonate, and combinations thereof.

7. The method of claim 1, wherein said sacrificial polymer is selected from poly[(oxycarbonyloxy-1,1,4,4-tetramethylbutane)-alt-(oxycarbonyloxy-p-xylene)], poly[(oxycarbonyloxy-1,4-dimethylbutane)-alt-(oxycarbonyloxy-p-xylene)], and combinations thereof.

\* \* \* \* \*